CHEM 12 Hess's Law Worksheet

 Calculate the ∆H for the reaction: C₂H₄ (g) + H₂ (g) → C₂H₆ (g), from the following data:

$$\begin{array}{ll} C_2H_4\ (g) + 3\ O_2\ (g) \Rightarrow 2\ CO_2\ (g) + 2\ H_2O\ (l) & \Delta H = -\ 1411\ kJ\ /\ mole \\ C_2H_6\ (g) + 7/2\ O_2\ (g) \Rightarrow 2\ CO_2\ (g) + 3\ H_2O\ (l) & \Delta H = -\ 1560.\ kJ\ /\ mole \\ H_2\ (g) + \frac{1}{2}\ O_2\ (g) \Rightarrow H_2O\ (l) & \Delta H = -\ 285.8\ kJ\ /\ mole \end{array}$$

2. Calculate the ΔH for the reaction:

$$\begin{array}{ll} N_2 \, (g) + O_2 \, (g) \rightarrow 2 \; NO \, (g) & \Delta H = -180.5 \; kJ \; / \; mole \\ N_2 \, (g) + 3H_2 \, (g) \rightarrow 2 \; NH_3 \, (g) & \Delta H = -91.8 \; kJ \; / \; mole \\ 2 \, H_2 \, (g) + O_2 \, (g) \rightarrow 2 \; H_2O \, (g) & \Delta H = -483.6 \; kJ \; / \; mole \end{array}$$

3. Find the $\Delta H_{\rm f}$ ° (heat of formation) for acetic acid, $HC_2H_3O_2$, using the following thermochemical data:

$$HC_2H_3O_2(1) + 2 O_2(g) \rightarrow 2 CO_2(g) + 2 H_2O(1)$$
 $\Delta H = -875 \text{ kJ / mole}$ $C(s) + O_2(g) \rightarrow CO_2(g)$ $\Delta H = -394.51 \text{ kJ / mole}$

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