

* Show that, the entropy remains constant during a reversible cyclic change.

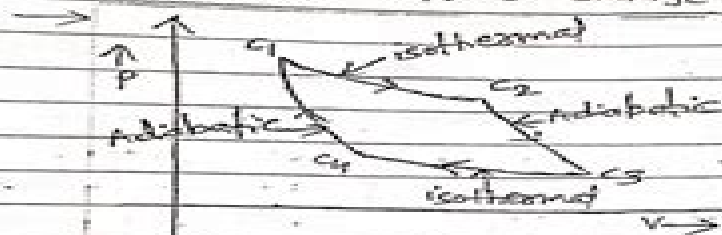


Fig: Carnot's cycle

Considered a reversible cycle C_1, C_2, C_3, C_4, C_1 as shown in fig. It consists of 4 operations
 1) From C_1 to C_2 heat Q_1 is absorbed by the system at temp. T_1 . The change in entropy of the system is S_1 . The change is positive.

2) From C_2 to C_3 , no heat is absorbed or rejected by the system, hence no change in entropy

3) From C_3 to C_4 , heat Q_2 is rejected by the system at temp. T_2 . The change in entropy is S_2 . The change is negative.

4) From C_4 to C_1 , no heat is absorbed or rejected by the system, hence no change in entropy.

$\therefore$ Total change in entropy of a working substance during reversible cycle is $\frac{Q_1}{T_1} = \frac{Q_2}{T_2} \Rightarrow \frac{Q_1}{T_1} - \frac{Q_2}{T_2} = 0$

change in entropy of the system of source & sink is zero. Thus during reversible cycle, the entropy of the universe remains constant. i.e. $\Delta S_R = 0$.

* state & explain the principle of increase of entropy or explain entropy of universe in irreversible processes.

Efficiency of Carnot's cycle is

$$\eta = 1 - \frac{Q_2}{Q_1}$$

But, this efficiency is less than that of a reversible engine working between the same two temp's T_1 & T_2 . for which, efficiency

$$\eta = 1 - \frac{T_2}{T_1}$$

$$\therefore 1 - \frac{Q_2}{Q_1} < 1 - \frac{T_2}{T_1}$$

Applied Thermodynamics Notes

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