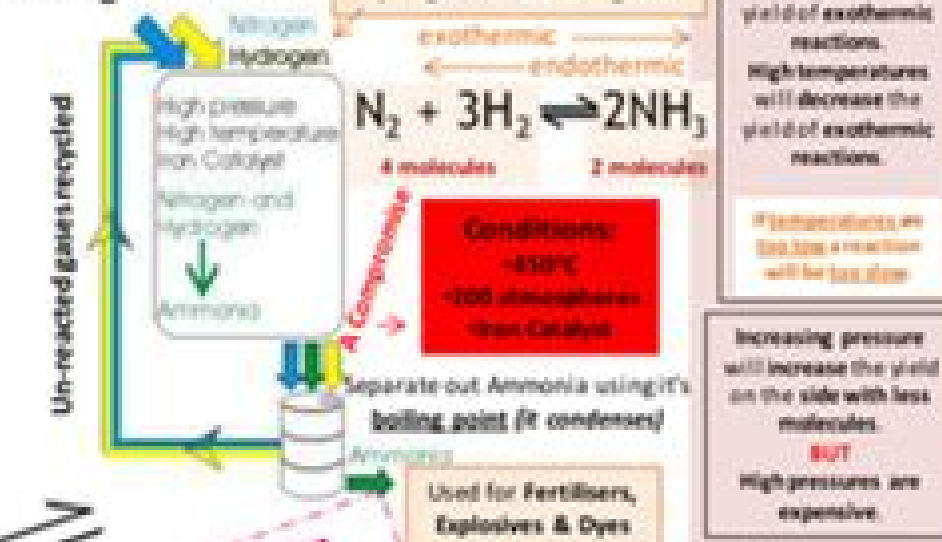


# The Haber Process Making Ammonia



**Reversible Reaction**  
 One way will be exothermic and the other will be endothermic.  
 In a closed system forward and backward reaction occur at equal rates.

## Flame Test

- Dip wire in sample  
- Hold loop in flame  
- Use the colour of the flame to identify the metal in the compound

Colour	Ion
Crimson / Bright Red	$\text{Li}^+$
Brilliant Red	$\text{Ca}^{2+}$
Yellow	$\text{Na}^+$
Lilac	$\text{K}^+$
Green	$\text{Ba}^{2+}$

## Negative Ions

1) Halide ions  $\text{Cl}^-$   $\text{Br}^-$   $\text{I}^-$  → Add Nitric acid → Add Silver Nitrate → Look for coloured precipitate

2) Sulfates ions  $\text{SO}_4^{2-}$  → Add acid → Add Barium chloride → Look for white precipitate

3) Carbonate ions  $\text{CO}_3^{2-}$  → Add acid → Look for bubbles and test w/ Litmus (acid is made)

## Reactions with NaOH

- Add NaOH to sample  
- Check to see what coloured precipitate (solid) forms.  
- Check to see if adding more NaOH will make it dissolve

Unknown solution + NaOH	Observation	Ion
White precipitate	White precipitate	$\text{Al}^{3+}$
Light blue precipitate	Light blue precipitate	$\text{Cu}^{2+}$
Light green precipitate that starts to go brown	Light green precipitate that starts to go brown	$\text{Fe}^{2+}$
Reddish brown precipitate	Reddish brown precipitate	$\text{Fe}^{3+}$
Went dissolve with more NaOH	Went dissolve with more NaOH	$\text{Mg}^{2+}$ $\text{Ca}^{2+}$

## Titration

Acids and alkalis neutralise each other. You can use an indicator to show when this has happened. This is the **END POINT**.

Repeat three times until you can make an average.

**Remember:**  
 Acid + Alkali → Salt + Water

**Chemical Equation:**  

$$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$$

## Acid Strength?

Two acids that are the same concentration can and be different strengths.

**Strong acids dissociate** (break) there in ions completely.  
**Weak acids do not completely dissociate**, not being of their full strength.

**Homologous series** - a group of compounds which share a functional group giving them similar properties.

**All alcohols have the functional group -OH**

**All carboxylic acids have the functional group -COOH**

**All esters have the functional group -COO-**

## Calculation

Known Acid	Conc. of Unknown Alkali
Concentration (mol/dm³)	0.5
Volume (dm³)	25/1000 = 0.025
Moles (n)	0.5 x 0.025 = 0.0125

**Remember:**  
 A titration uses **known** of a known acid (0.5 mol/dm³) to neutralise **known** of an unknown alkali.

**First work out moles in known acid.**  
 Use equation to work out ratio of moles and apply.  
 (Moles 1 : 1 reaction used as the number of moles is the same)  
 Next use the volume and moles to calculate the concentration of unknown alkali.  
 The unknown alkali has a concentration of 1 mol/dm³

## Forensic gel electrophoresis

Chemical analysis can be used to ID individuals at crime scenes, pollution in rivers and matching transplant patients to donors in medicine.

Crime scene	Prison	Donor
1	2	3
4	5	6
7	8	9
10	11	12
13	14	15
16	17	18
19	20	21
22	23	24
25	26	27
28	29	30
31	32	33
34	35	36
37	38	39
40	41	42
43	44	45
46	47	48
49	50	51
52	53	54
55	56	57
58	59	60
61	62	63
64	65	66
67	68	69
70	71	72
73	74	75
76	77	78
79	80	81
82	83	84
85	86	87
88	89	90
91	92	93
94	95	96
97	98	99
100	101	102

The time to test these samples is subject to change for technical reasons

# Aqa Chemistry C3 January 2013 Paper

**Terry C. Jones**



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