Name:			

Date: _____

Period:

S'mores Stoichiometry Lab

In troduction:

In this experiment, you will explore the principles of stoichiometry by building S'mores, the delicious, chocolate, marshmallow, and graham cracker treats.



Definitions. 'Stoichio' means element and 'metry' means the process of measuring. The mess and quantity relationships among reacters and products in a reaction are found using the process of stoichiometry.

Purpose:

To determine the limiting reactent in the synthesis of S'mores.

Prelib Questions:

- What does a chemical equation tell us?
- What do the coefficients in a belenced chemical equation represent?

Materials:

Ministure Chocolate bars Marshmallows Buosen Burner Graham cracker Square Paper Towels Wooden skewers

Reference Data:

Substance	Symbol	Unit mass	Package Mass
Graham Cracker Squares	Gr	8.50 g	403 g
Marshmallow	Mm	7.125 g	285 g
Miniature Chocolate Bar	Ch	12.00 g	348 g

** for the purposes of this lab the UNIT mass and MOLAR mass are the same thing

Procedure:

- Perform a synthesis reaction (forming one s'more). For this lab, we will consider a s'more to consist of 2 graham cracker squares, I marshmallow and I ministure chocolate bar.
 - a) Determine the formula for a s'more
 - write the balanced chemical equation using the symbols above
 - c) According to the belanced equation complete the following ratios:

Or -	Man	G: =	S'es ore
Ch-	Mm	Mm =	S"su ore
Ch-	Gr	Ch =	S'more

- Calculating the unit mass of the s'more.
 - a) Calculate the mass on ONE s'more.
 - b) Is there a relationship between the mass of a S'more and the masses of the reactants used to make it?
 - c) If so, what is the relationship?
 - d) What law have you studied in this course that might define this relationship?
- 3. Predicting the limiting reactant
 - a) Complete the data table below based on the materials provided for your group.

Substance	Quantity received	# of S'm cres it COULD make
Graham Cracker Squares		
Marshmallow		
Miniature Chocolate Bar		

- b) What will the limiting reactant be?
- c) How many of the other substances will remain
- d) What are the masses of each excess reactant that will remain once the reaction is complete?

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EW Minium

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