

16.1**PROPERTIES OF SOLUTIONS****Section Review****Objectives**

- Identify the factors that determine the rate at which a solute dissolves
- Identify the units usually used to express the solubility of a solute
- Calculate the solubility of a gas in a liquid under various pressure conditions
- Identify the factors that determine the mass of solute that will dissolve in a given mass of a solvent

Vocabulary

- saturated solution
- solubility
- unsaturated solution
- miscible
- immiscible
- supersaturated solution
- Henry's law

Key Equation

- Henry's law: $\frac{S_1}{P_1} = \frac{S_2}{P_2}$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Changes in the temperature of a system and 1 of a solute alter the 2 at which a solute dissolves. The extent to which a gas dissolves in a liquid is proportional to the 3 of the gas in accordance with 4 law. The solubility of a gas decreases with increasing 5. A solution that contains the maximum amount of solute at a given temperature is said to be 6. Two liquids that are mutually soluble in each other are said to be 7. Generally the 8 of a solid in water 9 with increasing temperature, but there are exceptions. A(n) 10 solution holds more solute than is theoretically possible.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
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9. _____
10. _____

161 Properties Of Solutions Section Review Answers

Mary Millhollon, Richard H. Langley



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