

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014  
(Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	$1.6725 \times 10^{-24}$	$1.6725 \times 10^{-24}$	$0.0009 \times 10^{-24}$

- i An atom of hydrogen can be represented as  ${}^1\text{H}$ . Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

- ii Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A electron  
B neutron  
C proton

Write the correct letter, A, B, or C, on the line below.

B

(1 mark)

- b A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes,  ${}^{10}\text{B}$  and  ${}^{11}\text{B}$ .

- i Calculate the percentage abundance of  ${}^{10}\text{B}$  in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

# Aqa A2 Chemistry Exam Style Questions Answers

**Michael Scaife**



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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Aqa A2 Chemistry Exam Style Questions Answers**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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