

Name: \_\_\_\_\_ Per: \_\_\_\_\_  
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Honors Chemistry-Smith

Average Atomic Mass Practice Problems

The answers are given in bold and italics after each problem. To receive credit, you must show work. The work must be legible and organized.

- 1) Calculate the average atomic mass of lithium to the nearest thousandth, which occurs as two isotopes that have the following atomic masses and abundances in nature: 7.30% at 6.017 amu, and 92.70% at 7.018 amu. (*6.945 amu*)

$$7.3\% \times 6.017 + 92.7\% \times 7.018 = 7.028 \text{ amu}$$

- 2) Calculate the average atomic mass of magnesium using the following data for three magnesium isotopes to 3 decimal places. (*24.447 amu*)

Isotope	mass (amu)		relative abundance
Mg-24	23.985	X	0.7870 <del>+</del>
Mg-25	24.986	X	0.1013 <del>+</del>
Mg-26	25.983	X	0.1117

$$= 24.310 \text{ amu}$$

- 3) Lithium has two naturally occurring isotopes: lithium-6 and lithium-7. If the average atomic mass of lithium is 6.941 amu, which isotope is the most abundant? How do you know?

*Lithium - 7  
 6.941 is closer to 7.*

- 4) The average atomic mass of copper is 63.55 amu. If the only two isotopes of copper have masses of 62.94 amu and 64.93 amu, what are the percentages of each (2 decimal places)? (This involves algebra) (*69.35% of 62.94, 30.65% of 64.93*)

*Let the percentage of 62.94 amu be x.*

$$62.94x + 64.93(1-x) = 63.55 \quad x = 0.6935 = 69.35\%$$

$$62.94x + 64.93 - 64.93x = 63.55 \quad 1 - 0.6935 = 30.65\%$$

$$-1.99x = -1.38$$

- 5) Naturally occurring silicon consists of three isotopes,  $^{28}\text{Si}$ ,  $^{29}\text{Si}$ , and  $^{30}\text{Si}$ , whose atomic masses are 27.9769, 28.9865, and 29.9838, respectively. The most abundant isotope is  $^{28}\text{Si}$ , which accounts for 92.23% of naturally occurring silicon. Given that the observed atomic mass of silicon is 28.0855, calculate the percentages of  $^{29}\text{Si}$  and  $^{30}\text{Si}$  in nature. (*4.75% of  $^{29}\text{Si}$ , 3.02% of  $^{30}\text{Si}$* )

$$27.9769 \times 0.9223 + 28.9865x + 29.9838(0.0777 - x) = 28.0855$$

$$x = 4.75\% \quad 0.0777 - 0.0475 = 3.02\%$$

# Chemistry Practice Problems Answers Atomic Mass

**Wolfgang Guggemos**



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## **Table of Contents Chemistry Practice Problems Answers Atomic Mass**

1. Understanding the eBook Chemistry Practice Problems Answers Atomic Mass
  - The Rise of Digital Reading Chemistry Practice Problems Answers Atomic Mass
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Problems Answers Atomic Mass
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Practice Problems Answers Atomic Mass
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Problems Answers Atomic Mass
  - Personalized Recommendations
  - Chemistry Practice Problems Answers Atomic Mass User Reviews and Ratings
  - Chemistry Practice Problems Answers Atomic Mass and Bestseller Lists
5. Accessing Chemistry Practice Problems Answers Atomic Mass Free and Paid eBooks
  - Chemistry Practice Problems Answers Atomic Mass Public Domain eBooks
  - Chemistry Practice Problems Answers Atomic Mass eBook Subscription Services
  - Chemistry Practice Problems Answers Atomic Mass Budget-Friendly Options

6. Navigating Chemistry Practice Problems Answers Atomic Mass eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Practice Problems Answers Atomic Mass Compatibility with Devices
  - Chemistry Practice Problems Answers Atomic Mass Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Practice Problems Answers Atomic Mass
  - Highlighting and Note-Taking Chemistry Practice Problems Answers Atomic Mass
  - Interactive Elements Chemistry Practice Problems Answers Atomic Mass
8. Staying Engaged with Chemistry Practice Problems Answers Atomic Mass
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Practice Problems Answers Atomic Mass
9. Balancing eBooks and Physical Books Chemistry Practice Problems Answers Atomic Mass
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Practice Problems Answers Atomic Mass
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practice Problems Answers Atomic Mass
  - Setting Reading Goals Chemistry Practice Problems Answers Atomic Mass
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Practice Problems Answers Atomic Mass
  - Fact-Checking eBook Content of Chemistry Practice Problems Answers Atomic Mass
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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