

Class 11 Redox Reactions Important Questions with Answers

Short Answer Type Questions

Q1. The reaction



represents the process of bleaching. Identify and name the species that bleaches the substances due to their oxidising action.

Answer:



Here, Cl_2 is both oxidized and reduced in ClO^- and Cl^- , respectively. Since Cl^- cannot act as an oxidising agent (O.A.). Therefore, Cl_2 bleaches substances due to oxidising action of hypochlorite ClO^- ion.

Q2. MnO_4^{2-} undergoes a disproportionation reaction in an acidic medium but MnO_4^- does not. Give a reason.

Answer:

Disproportionation is a redox reaction in which one intermediate oxidation state component transforms into two higher and lower oxidation state compounds.

Manganese's oxidation states ranging from +2 to +7 in its various compounds. MnO_4^- has the maximum oxidation state of +7 hence disproportionation is impossible, but MnO_4^{2-} has a +6 oxidation state, which can be oxidised as well as reduced.

Q3. PbO and PbO_2 react with HCl according to following chemical equations:



Why do these compounds differ in their reactivity?

Answer:

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M Walker



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