Reaction Rates and Mechanisms Quiz

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1. Within a reaction with 2 machini and 2 products, the real tion rate can be averaged with respect to:

At all 4 adictances B13 of the adictances C12 of the adictances

Di only I of the rebritance: E) not enough information to determine

2. The minimum energy required for a reaction to take place is called the

A) Activition energy (ii) Activited complex (ii) Newtof mection (iii) Enthalpy

2. A temperary structure with a high potential energy that has bonds half formed and half broken is:

a) a homogeneous cutalyst b) an activated complex c) milital

d) an elementary process | e| a beterogeneous cutalyst

4. Consider the following mactions at 25°C.

LAc Yest + Bir but -- Activit III 4Fe(st + 300stc) -- 2Fe00sts

Which of the following statements best describes the relative rates of the two reactions?

A. I and III are both fact. C. I is factor than II.

E. I and III are both sizes. D. II is factor than II.

5. What property would be appropriate to measure the reaction rate in the following reaction:

 $\alpha M_{\rm BH} + \beta \Omega a B O_{\rm BHH} \longrightarrow M_{\rm A} \Omega B O_{\rm A} S_{\rm BHH} + \beta \Omega S_{\rm BHH}$

a change in conductivity by change in maco or change in colour

d charge in pressure e, charge in volume

6. The theoretical effect of an increase in the initial concentration of a reaction can be explained in terms of collision theory because it affects:

the collision peometry involved in the reaction.

II. the total journiber of collisions that occur

III. the fraction of collisions that are effective

TV. the required activistics energy for a reaction.

At both I and IV Fold III and III are true CI both III and IV Di both II and III E) II only

At a certain temperature, 2.50 g Cu reacts completely in 20.0 seconds.
 The rate of consumption of Cu is:

 $Cr_{\rm col} + 2H_2O_{\rm col} \rightarrow Cal(OU)_{\rm bard} + H_{\rm Spel}$

A) 0.00200 mol/min B) 0.0000 mol/min C) 0.125 mol/min D) 5.00 mol/min

Chemistry Quiz Rate Of Reaction Answer Key

Clifford Lane

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