

### Worksheet #3—Stoichiometry

**Directions:** Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation:  $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

|               |                           |  |                                 |
|---------------|---------------------------|--|---------------------------------|
| 14.8 moles Al | 3 moles of O <sub>2</sub> |  | = <u>11.1 mol O<sub>2</sub></u> |
|               | 4 mole Al                 |  |                                 |

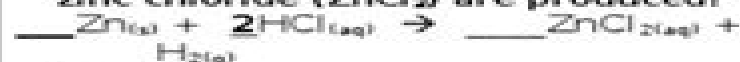
2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO<sub>3</sub>).



|                              |                           |  |   |
|------------------------------|---------------------------|--|---|
| 0.61 moles KClO <sub>3</sub> | 3 moles of O <sub>2</sub> |  | = |
|                              | 2 mole KClO <sub>3</sub>  |  |   |

0.915 mol O<sub>2</sub>

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl<sub>2</sub>) are produced?



|         |          |                          |                         |                                  |
|---------|----------|--------------------------|-------------------------|----------------------------------|
| 20 g Zn | 1 mol Zn | 1 mole ZnCl <sub>2</sub> | 136 g ZnCl <sub>2</sub> |                                  |
|         | 65g Zn   | 1 mol Zn                 | 1 mol ZnCl <sub>2</sub> | = <u>41.8 g ZnCl<sub>2</sub></u> |

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



|           |            |                       |                       |                                |
|-----------|------------|-----------------------|-----------------------|--------------------------------|
| 10 g NaCl | 1 mol NaCl | 1 mol Cl <sub>2</sub> | 71 g Cl <sub>2</sub>  |                                |
|           | 58.5g NaCl | 2 mol NaCl            | 1 mol Cl <sub>2</sub> | = <u>6.07 g Cl<sub>2</sub></u> |

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



|                        |                         |                         |           |  |
|------------------------|-------------------------|-------------------------|-----------|--|
| 4.0g AgNO <sub>3</sub> | 1 mol AgNO <sub>3</sub> | 1 mole Cu               | 63.5 g Cu |  |
|                        | 170g AgNO <sub>3</sub>  | 2 mol AgNO <sub>3</sub> | 1 mol Cu  |  |

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



|                        |                       |                        |  |
|------------------------|-----------------------|------------------------|--|
| 15.1 g Cl <sub>2</sub> | 1 mol Cl <sub>2</sub> | 1 mole H <sub>2</sub>  |  |
|                        | 71 g Cl <sub>2</sub>  | 1 mole Cl <sub>2</sub> |  |

= 0.21 moles H<sub>2</sub>

# Chemistry Stoichiometry Problems And Answers

**K.V. NARAYANAN ,B. LAKSHMIKUTTY**



## **Chemistry Stoichiometry Problems And Answers:**

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With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations The book is supplemented with Solutions Manual for instructors containing detailed solutions of all chapter end unsolved problems NEW TO THE SECOND EDITION Incorporates a new chapter on Bypass Recycle and Purge Operations Comprises updations in some sections and presents new sections on Future Avenues and Opportunities in Chemical Engineering Processes in Biological and Energy Systems Contains several new worked out examples in the chapter on Material Balance with Chemical Reaction Includes GATE questions with answers up to the year 2016 in Objective type questions KEY FEATURES SI units are used throughout the book All basic chemical engineering operations and processes are introduced and different types of problems are illustrated with worked out examples Stoichiometric principles are extended to solve problems related to bioprocessing environmental engineering etc Exercise problems more than 810 are organised according to the difficulty level and all are provided with answers

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*Multiple Solution Methods for Teaching Science in the Classroom* Stephen DeMeo, 2008 For the first time in science education the subject of multiple solution methods is explored in book form While a multiple method teaching approach is utilized extensively in math education there are very few journal articles and no texts written on this topic in science Teaching multiple methods to science students in order to solve quantitative word problems is important for two reasons First it challenges the practice by teachers that one specific method should be used when solving problems Secondly it calls into question the belief that multiple methods would confuse students and retard their learning Using a case study approach and informed by research conducted by the author this book claims that providing students with a choice of methods as well as requiring additional methods as a way to validate results can be beneficial to student learning A close reading of the literature reveals that time spent on elucidating concepts rather than on algorithmic methodologies is a critical issue when trying to have students solve problems with understanding It is argued that conceptual understanding can be enhanced through the use of multiple methods in an environment where students can compare evaluate and verbally discuss competing methodologies through the facilitation of the instructor This book focuses on two very useful methods proportional reasoning PR and dimensional analysis DA These two methods are important because they can be used to solve a large number of problems in all of the four academic sciences biology chemistry physics and earth science This book concludes with a plan to integrate DA and PR into the academic science curriculum starting in late elementary school through to the introductory college level A challenge is presented to teachers as well as to textbook writers who rely on the single method paradigm to consider an alternative way to teach scientific problem solving

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**Dr. Salm's Chemistry Problem Drill Book** Robert J. Salm, Robert J. Salm, Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College. This work will give you practice with the major problem types as you prepare for finals and standardized tests.

**Chemistry** Therald Moeller, 2012-12-02 *Chemistry with Inorganic Qualitative Analysis* is a textbook that describes the application of the principles of equilibrium represented in qualitative analysis and the properties of ions arising from the reactions of the analysis. This book reviews the chemistry of inorganic substances as the science of matter, the units of measure used, atoms, atomic structure, thermochemistry, nuclear chemistry, molecules and ions in action. This text also describes the chemical bonds, the representative elements, the changes of state, water and the hydrosphere, which also covers water pollution and water purification. Water purification occurs in nature through the usual water cycle and by the action of microorganisms. The air flushes dissolved gases and volatile pollutants when water seeps through the soil; it filters solids as they settle in the bottom of placid lakes. Microorganisms break down large organic molecules containing mostly carbon, hydrogen, nitrogen, oxygen, sulfur, or phosphorus into harmless molecules and ions. This text notes that natural purification occurs if the level of contaminants is not so excessive. This textbook is suitable for both chemistry teachers and students.

**Resources in Education**, 1994      **A Visual Analogy Guide to Chemistry, 2e** Paul A Krieger, 2018-02-01 *A Visual Analogy Guide to Chemistry* is the latest in the innovative and widely used series of books by Paul Krieger. This study

guide delivers a big picture view of difficult concepts and effective study tools to help students learn and understand the details of general organic and biochemistry topics A Visual Analogy Guide to Chemistry is a worthwhile investment for any introductory chemistry student

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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