

ChemThink: Covalent Bonding

Covalent bonding forms when atoms are sharing electrons.

When two atoms get close enough, the nucleus attracts the other atom's
 1. Electrons 2. Neutrons 3. Protons

Before bonding, the atom's electrons spend most of their time
orbiting the nuclei of each atom. Once bonded, the
 electrons spend most of their time between the two nuclei.

Atoms must be able to hold onto their own electrons while
pulling attraction other atom's electrons.

Covalent bonds form between two nonmetals.

When atoms move closer, the potential energy increases decreases.
 At a certain point the potential energy increases. If you try
 and force the atoms closer against the protons in each
 nucleus are repelling each other.

The ideal distance between the atoms is known as the bond length.

Lower in energy = more stable.

| Bond Type | Draw an Example | # of paired e ⁻ | Total # of e ⁻ shared | Strongest Weakest |
|-----------|-----------------|----------------------------|----------------------------------|-------------------|
| Single | $H-H$ | 1 | 2 | <u>weakest</u> |
| Double | $O=O$ | 2 | 4 | — |
| Triple | $N \equiv N$ | 3 | 6 | <u>strongest</u> |

Chemthink Answers Covalent Bonding Molecular Shape

Marcel A. Müller



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