

- consider this reaction (the capital letters are the chemical, and the lower case letters are the balancing coefficient):

$$aA + bB \rightarrow cC + dD$$
 - The equilibrium constant (K_{eq}) is the *ratio of product concentration to the reactant concentration at equilibrium*, with each concentration *raised to a power* (which is the balancing coefficient).
- consider this reaction:

$$aA + bB \rightarrow cC + dD$$
 - Thus, the “**equilibrium constant expression**” has this general form:

(brackets: [] = molarity concentration)

- the equilibrium constants provide valuable information, such as whether products or reactants are favored:
 if $K_{eq} > 1$, _____ favored at equilibrium
 if $K_{eq} < 1$, _____ favored at equilibrium

Sample Problem: Expressing and Calculating K_{eq}

The colorless gas dinitrogen tetroxide (N_2O_4) and the dark brown gas nitrogen dioxide (NO_2) exist in equilibrium with each other.



A liter of gas mixture at equilibrium at 10 °C contains 0.0045 mol of N_2O_4 and 0.030 mol of NO_2 . Write the expression for the equilibrium constant and calculate the equilibrium constant (K_{eq}) for the reaction

Chp 18 Rates Of Reaction Packet

Mark Steyvers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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