

## Chp 8 Study guide:

Using a T-chart on a separate piece of binder paper, write the question and then answer the question.

- 1) A chemical equation like  $\text{Na} + \text{Cl} \rightarrow \text{NaCl}$  proves what law?
- 2) If you have 50 grams of a reactant you will end up with 50 grams of a product. What law is this?
- 3) How does the following equation break the law of conservation of matter?  
 $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 4) Use the following equation to answer the questions below  
 $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ 
  - a) How many products produced?
  - b) How many reactants are produced?
- 5) For an exothermic reaction, the heat content of the \_\_\_\_\_ is greater than that of the \_\_\_\_\_. Draw a diagram to support this.
- 6) For an endothermic reaction, the heat content of the \_\_\_\_\_ is greater than that of the \_\_\_\_\_. Draw a diagram to support this.
- 7) When two or more elements react chemically what can you say about their chemical and physical properties in comparison to the original element.
- 8) How can elements form compounds/molecules? Explain
  - a) Chemical reactions
  - b) Physical changes
  - c) State of changes

- 9) List the types of evidence that will indicate a chemical change.
- 10) If one of your reactants weighs 50 grams and your product weighs 100 grams, what is the mass of your other reactants? What law is this?
- 11) Give three examples of a chemical change.
- 12) Give three examples of a physical change.
- 13) Give two examples of an endothermic reaction.
- 14) Give two examples of an exothermic reaction.
- 15) What type of change has occurred in the below equation, a chemical or physical change? Explain



Balance the following equations

- 16) \_\_\_\_  $\text{H}_3\text{PO}_4 \rightarrow$  \_\_\_\_  $\text{H}_4\text{P}_2\text{O}_7$  + \_\_\_\_  $\text{H}_2\text{O}$
- 17) \_\_\_\_  $\text{KClO}_3 \rightarrow$  \_\_\_\_  $\text{KCl}$  + \_\_\_\_  $\text{O}_2$
- 18) \_\_\_\_  $\text{H}_2\text{CO}_3 \rightarrow$  \_\_\_\_  $\text{H}_2\text{O}$  + \_\_\_\_  $\text{CO}_2$

# Chp 8 Study Guide Answers Houghton Mifflin

**M Woodhall**



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