

CHEMISTRY Y12
The Acidic Environment
CLOZE PASSAGE No 4
The History of Acids

Complete the following sentences using appropriate words or short phrases

a) Chemist who proposed that acids were oxygen containing substances	Lavoisier
b) Humphry Davy proposed that was the element present in all acids	Hydrogen
c) Arrhenius theory states "An acid is a substance that in water to produce ions as the onlyions in solution"	Dissolves,Hydrogen, Positive
d) A water soluble base should contain ions	Hydroxide OH ⁻
e) Name the 2 chemists that poposed an acid is a proton donor	Lowry, Bronsted
f) Using the LB model complete the equation $\text{HCl} + \text{H}_2\text{O} \rightarrow$	$\text{H}_3\text{O}^+ + \text{Cl}^-$
g) State the conjugate acid for CO_3^{2-}	HCO_3^-
h) State the conjugate base for NH_4^+	NH_3
i) A species that can be either an acid or base is said to be	Amphiprotic
j) Identify if the following salts when added to water are Acid, Neutral or Bases Sodium Chloride, Ammonium Chloride, Silver Acetate	Neutral,Acid,Base
k) The reaction of any species with water is called	Hydrolysis
l) The colour of universal indicator in Sodium Carbonate	Purple it is basic
m) The name of the domestic acid substance you used to determined its concentration	Vinegar
n) 2 pieces of calibrated glassware is in titrations	Pipette,Burette
o) Used to rinse the conical flask in a titration	Distilled Water
p) The chemical used for our primary standard	Sodium carbonate
q) The point at which the indicator changes colour	End Point
r) Give 2 examples of amphiprotic species	H_2O and HCO_3^-
s) A suitable indicator for determining the conc of acetic acid in vinegar	Phenolphthalein
t) A buffer is a mixture of a w..... acid and its c..... base	Weak , Conjugate
u) An example of a naturally buffered system	Blood/Seawater
v) 2 chemicals that could be added to dilute and neutralise an acid spill	Water, Sodium Carbonate
w) A computer based technology used in titration experiments	Data logger attached to a pH probe
x) The name given to the volume of liquid delivered from a burette	titre
y) The name of the flask used to prepare a primary standard	Volumetric
z) In an ideal titration the point and end point coincide	Equivalence

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Stoichiometry

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