

## Periodic Trends

Can the properties of an element be predicted using a periodic table?

### Why?

The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

3 1. Consider the data in Model 1 on the following page.

a. Each element has three numbers listed under it. Which value represents the atomic radius?

The first number below the diagram

b. What are the units for the atomic radius?

pm picometers  $10^{-12}$  m

c. Write a complete sentence to convey your understanding of atomic radius. *Note:* You may not use the word "radius" in your definition.

Distance from center of atom (nucleus) to outer electrons

2 2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.

Increases  $H < Li < Na < K$   $F < Cl < Br$   
 $C < Si < Ge$

3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. *Hint:* You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.

As the number of electrons increases, new energy levels need to be added further from the nucleus, increasing the radius.

2 4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.

Decreases  $Li > Be > B > C > N > O > F > Ne$   
 $Na > Mg > Al > Si > P > S > Cl > Ar$

5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.

Highest energy level is constant, but the number of protons in the nucleus increases, increasing the charge of the nucleus and the attractive force on the electrons, pulling the highest energy level closer to the nucleus.

# Chemistry Periodic Trends Answer Key

**Arshad Iqbal**



## **Chemistry Periodic Trends Answer Key:**

Periodic Table Cindy Blobaum, 2005 What do chocolate chip cookies chemistry and logic have in common They are the basis for a unit that lets students become actively engaged in discovering the arrangement of the periodic table This learning activity takes the periodic table out of the static presentation usually associated with textbooks and chemistry courses and interjects an element of discovery The two activities in this unit provide students with information that they have to arrange in organized charts In the process of creating the arrangements students will be involved in problem solving and will gain an appreciation for the scientific process of exploration and verification This dynamic unit meets national science standards in seven teaching and content areas Bring the periodic table to life with this hands on minds on unit Book jacket **Kaplan SAT Subject Test Chemistry 2015-2016** Kaplan Test Prep, 2015-03-03 Essential strategies practice and review to ace the SAT Subject Test Chemistry Getting into a top college has never been more difficult Students need to distinguish themselves from the crowd and scoring well on a SAT Subject Test gives students a competitive edge Kaplan s SAT Subject Test Chemistry is the most up to date guide on the market with complete coverage of both the content review and strategies students need for success on test day Kaplan s SAT Subject Test Chemistry features A full length diagnostic test Full length practice tests Focused chapter summaries highlights and quizzes Detailed answer explanations Proven score raising strategies End of chapter quizzes Kaplan is serious about raising students scores we guarantee students will get a higher score **Chemistry for the Utterly Confused** John T. Moore, Richard H. Langley, 2007-06-01 Banish bafflement in this tough subject From formulas and lab techniques to the periodic table Chemistry for the Utterly Confused focuses on the areas of maximum confusion and breaks down the most difficult chemistry topics into easy to understand concepts This invaluable guide also teaches problem solving skills you need to master this imposing subject Whether you re in high school in college or simply brushing up on chemistry knowledge this fun easily accessible book will make understanding chemistry a breeze **GO TO Objective NEET 2021 Chemistry Guide 8th Edition** Disha Experts, **NEET Chemistry 1500+ MCQs** Disha Experts, 2019-12-24 *JEE Main Chapter Wise Numerical Response Questions with Solution for Chemistry (As Per NTA Latest Pattern )* Career Point Kota, 2020-07-21 Whenever a student decides to prepare for any examination her his first and foremost curiosity is about the type of questions that he she has to face We feel great pleasure to present this book before you We have made an attempt to provide Chapter wise Numerical Response Questions for JEE Main as per NTA latest pattern with answer and solutions to majority of questions Solutions to the questions are not just sketch rather have been written in such a manner that the students will be able to understand the application of concept and can answer some other related questions too We firmly believe that the book in this form will definitely help a genuine hardworking student We have tried our best to keep errors out of this book Comment and criticism from readers will be highly appreciated and incorporated in the subsequent edition We wish to utilize the opportunity to place on record our special thanks to all team

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