

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base reaction |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is _____, a strong hydrogen ion donor?
2. _____ is the concentration of H^+ in an acid solution.
3. _____ is the concentration of OH^- in a basic solution.
4. _____ is the concentration of H^+ in an acid solution.
5. _____ is the concentration of OH^- in a basic solution.
6. _____ is the concentration of H^+ in an acid solution.
7. _____ is the concentration of OH^- in a basic solution.
8. _____ is the concentration of H^+ in an acid solution.
9. _____ is the concentration of OH^- in a basic solution.
10. _____ is the concentration of H^+ in an acid solution.
11. _____ is the concentration of OH^- in a basic solution.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____

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J Elliott



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