

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the quantity indicated that are indicated in the problem. Be sure to include the correct units and significant figures.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
3. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
4. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
5. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
6. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
7. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
8. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
9. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
10. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
11. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

7. _____

8. _____

9. _____

10. _____

11. _____

19 2 Hydrogen Ions And Acidity Worksheet Answers

DJ Losen



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