

# 18.2

## CONCENTRATIONS OF SOLUTIONS

### SECTION REVIEW

#### Objectives

- Solve problems involving the molarity of a solution.
- Describe how to prepare dilute solutions from more concentrated solutions of known molarity.
- Explain what is meant by percent by volume (PbV%) and percent by mass (PbM%) solutions.

#### Key Terms

- concentrated solution
- dilute solution
- concentrated solution
- molarity (M)

#### Key Equations

- Molarity (M) =  $\frac{\text{moles of solute}}{\text{liters of solution}}$
- $M_1 V_1 = M_2 V_2$
- Percent by volume (PbV%) =  $\frac{\text{volume of solute}}{\text{solution volume}} \times 100\%$
- Percent by mass (PbM%) =  $\frac{\text{mass of solute (g)}}{\text{solution volume (mL)}} \times 100\%$

#### Part A Completion

Use this completion activity to check your understanding of the concepts and terms discussed in this section. Each blank uses the completed word or term, alone, phrase, or sentence.

- The relative amounts of solute and \_\_\_\_\_ in a \_\_\_\_\_ can be described quantitatively as \_\_\_\_\_ or concentrated. Quantitative solutions of known concentration include molar concentrations, percent by volume, and percent by mass (volume).
- Molarity, the most important unit of concentration in chemistry, is expressed as \_\_\_\_\_ of solute per \_\_\_\_\_ of solution.
- Solutions of different concentrations can be prepared by \_\_\_\_\_ a stock solution. In dilution, the moles of \_\_\_\_\_ remain the same, while the amount of \_\_\_\_\_ changes.

# 182 Concentrations Of Solutions Section Review Answer Key

**Allan Blackman, Daniel  
Southam, Gwendolyn Lawrie, Natalie  
Williamson, Christopher Thompson**

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