- According to the Bronsted—Lowry definition, which chemical species can function both as an acid and as a base?
 - a. CI = 6. SO₂ =
 - SO/**
 NH/*
 - d. HCO₂
 - e. H₂O*
- 2. In the reaction CN⁻ + H₂O ← → HCN + OH⁻ which is an acid—base conjugate pair?
 - a. H₂O and HCN
 - b. H₂O and OH⁻
 - c. ON and H₂O
 - d. HCN and OH"
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX
 - a. It is less basic than a I M solution of NaY.
 - b. It is more basic than a I M solution of NaY.
 - It yields a neutral solution.
 - It is more concentrated than a 1 M solution of NaY.
- 4. The pH of a 0.03 M HCI solution is
 - a. 1.5
 - 6. 25
 - c. 3.5
 - d. 12.5
- The pH of a 1.0 x 10⁻³ M Ba(OH)₂ solution at 25°C is
 - a. 2.7
 - b. 3.0
 - c. 11.0
 - d. 11.3
- 6. Which is the strongest acid?
 - a. HCIO
 - A. HCIO₂
 - c. HOIO₃
 - d. HClO₄
- 7. Which salt reacts with water (hydrolyzes) to produce a basic solution?
 - a. NaC₂H₃O₂
 - NaNO₃
 - c. NH_fCI
 - d. BaSO₄

8. The weakest of the bases listed is

Acid	Conjugate Base	Ka (Ionization Constant of Acid)				
HCL	CIT	100% ionized				
HSO ₄ T	SO ₄ 2-	1.2 x 10°				
H ₂ S	HS -	5.7x10 ⁻⁴				
HS-	S'-	1.2x10 ⁻⁰				

- a. CIT
- # 1.00 T
- e. HS⁻ d. S³⁻
- e. 50e²⁻
- 9. The oxide of which element will react with water to form the strongest acid?

Period	Main Groups							
	1	п	П	IV	V	VI	VII	0
First		1						
Second	X	Y	Z	P		9	S	U
Third	WAY					R	T	0.40

- e. W
- 6. M
- c. P
- d. R.
- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate KC₂H₂O₂, is less alkaline than a 0.10 M solution of potassium evanide. KCN?
 - a. Hydrocyanic acid is a weaker acid than acetic acid.
 - Hydrocyanic acid is less soluble in water than acetic acid.
 - c. Cyanides are less soluble than acetaces
 - d. Acetic acid is a weaker acid than hydrocyanic acid
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of
 - a. a 0.100 M sodium solution.
 - a 0.0500 M sodium hydroxide solution.
 - c. a 0.0500 M benzoic acid solution.
 - a 0.0500 M sodium benzoate solution.

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- 1. Understanding the eBook Acs Inorganic Exam Practice Test
 - The Rise of Digital Reading Acs Inorganic Exam Practice Test
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Acs Inorganic Exam Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acs Inorganic Exam Practice Test
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Acs Inorganic Exam Practice Test
 - Personalized Recommendations
 - Acs Inorganic Exam Practice Test User Reviews and Ratings
 - Acs Inorganic Exam Practice Test and Bestseller Lists
- 5. Accessing Acs Inorganic Exam Practice Test Free and Paid eBooks

- Acs Inorganic Exam Practice Test Public Domain eBooks
- Acs Inorganic Exam Practice Test eBook Subscription Services
- Acs Inorganic Exam Practice Test Budget-Friendly Options
- 6. Navigating Acs Inorganic Exam Practice Test eBook Formats
 - o ePub, PDF, MOBI, and More
 - Acs Inorganic Exam Practice Test Compatibility with Devices
 - Acs Inorganic Exam Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acs Inorganic Exam Practice Test
 - Highlighting and Note-Taking Acs Inorganic Exam Practice Test
 - Interactive Elements Acs Inorganic Exam Practice Test
- 8. Staying Engaged with Acs Inorganic Exam Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acs Inorganic Exam Practice Test
- 9. Balancing eBooks and Physical Books Acs Inorganic Exam Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acs Inorganic Exam Practice Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acs Inorganic Exam Practice Test
 - Setting Reading Goals Acs Inorganic Exam Practice Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acs Inorganic Exam Practice Test
 - Fact-Checking eBook Content of Acs Inorganic Exam Practice Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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