

Once this is done, the mass of the gas can be computed from the amount and the molar mass:

$$m = (1.66 \times 10^2 \text{ mol}) \times (16.04 \text{ g mol}^{-1}) = 2.67 \times 10^3 \text{ g} = \boxed{2.67 \times 10^3 \text{ kg}}$$

E1.5(b) Use the equation for hydrostatic pressure [1.3], $p = p_{\text{ex}} + \rho gh$. Let p_{ex} be the pressure at the top of the straw and p is the pressure on the surface of the liquid (atmospheric pressure). Thus, the pressure difference is

$$\begin{aligned} p - p_{\text{ex}} &= \rho gh = (1.0 \text{ g cm}^{-3}) \times \frac{1 \text{ kg}}{10^3 \text{ g}} \times \left(\frac{1 \text{ cm}}{10^{-2} \text{ m}} \right)^3 \times (9.81 \text{ m s}^{-2}) \times (0.15 \text{ m}) \\ &= \boxed{1.5 \times 10^3 \text{ Pa}} = 1.5 \times 10^{-2} \text{ atm} \end{aligned}$$

E1.6(b) The pressure in the apparatus is given by

$$p = p_{\text{ex}} + \rho gh \quad [1.3]$$

where $p_{\text{ex}} = 760 \text{ Torr} = 1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$,

and $\rho gh = 13.55 \text{ g cm}^{-3} \times \left(\frac{1 \text{ kg}}{10^3 \text{ g}} \right) \times \left(\frac{1 \text{ cm}}{10^{-2} \text{ m}} \right)^3 \times 0.100 \text{ m} \times 9.806 \text{ m s}^{-2} = 1.33 \times 10^4 \text{ Pa}$

$$p = 1.013 \times 10^5 \text{ Pa} + 1.33 \times 10^4 \text{ Pa} = 1.146 \times 10^5 \text{ Pa} = \boxed{115 \text{ kPa}}$$

E1.7(b) Rearrange the perfect gas equation [1.8] to give $R = \frac{pV}{nT} = \frac{pV_{\text{m}}}{T}$

All gases are perfect in the limit of zero pressure. Therefore, the value of pV_{m}/T extrapolated to zero pressure will give the best value of R .

The molar mass can be introduced through

$$pV = nRT = \frac{m}{M} RT$$

which on rearrangement gives $M = \frac{m}{V} \frac{RT}{p} = \rho \frac{RT}{p}$.

The best value of M is obtained from an extrapolation of ρ/p versus p to zero pressure; the intercept is M/RT .

Draw up the following table:

p/atm	$(pV_{\text{m}}/T)/(\text{dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1})$	$(\rho/p)/(\text{g dm}^{-3} \text{ atm}^{-1})$
0.750 000	0.082 0014	1.428 59
0.500 000	0.082 0227	1.428 22
0.250 000	0.082 0414	1.427 90

From Figure 1.1(a), $\left(\frac{pV_{\text{m}}}{T} \right)_{p=0} = \boxed{0.082 062 \text{ dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1}}$

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Tobias Bleicker



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Instructor's Solutions Manual to Accompany Atkins' Physical Chemistry, Ninth Edition C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 The Instructor's solutions manual to accompany Atkins Physical Chemistry provides detailed solutions to the exercises and the even numbered discussion questions and problems that feature in the ninth edition of Atkins Physical Chemistry The manual is intended for instructors and consists of material that is not available to undergraduates The manual is free to all adopters of the main text

Physico-chemical Aspects of Textile Coloration Stephen M. Burkinshaw, 2016-02-08 The production of textile materials comprises a very large and complex global industry that utilises a diverse range of fibre types and creates a variety of textile products As the great majority of such products are coloured predominantly using aqueous dyeing processes the coloration of textiles is a large scale global business in which complex procedures are used to apply different types of dye to the various types of textile material The development of such dyeing processes is the result of substantial research activity undertaken over many decades into the physico chemical aspects of dye adsorption and the establishment of dyeing theory which seeks to describe the mechanism by which dyes interact with textile fibres *Physico Chemical Aspects of Textile Coloration* provides a comprehensive treatment of the physical chemistry involved in the dyeing of the major types of natural man made and synthetic fibres with the principal types of dye The book covers fundamental aspects of the physical and chemical structure of both fibres and dyes together with the structure and properties of water in relation to dyeing dyeing as an area of study as well as the terminology employed in dyeing technology and science contemporary views of intermolecular forces and the nature of the interactions that can occur between dyes and fibres at a molecular level fundamental principles involved in dyeing theory as represented by the thermodynamics and kinetics of dye sorption detailed accounts of the mechanism of dyeing that applies to cotton and other cellulosic fibres polyester polyamide wool polyacrylonitrile and silk fibres non aqueous dyeing as represented by the use of air organic solvents and supercritical CO₂ fluid as alternatives to water as application medium The up to date text is supported by a large number of tables figures and illustrations as well as footnotes and widespread use of references to published work The book is essential reading for students teachers researchers and professionals involved in textile coloration

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Science Reporter ,2003 The Rare Earth Trifluorides Boris Pavlovich Sobolev,2000 **Solid State Chemistry** Aaron Wold,Kirby Dwight,2012-12-06 The subject matter of solid state chemistry lies within the spheres of both physical and inorganic chemistry In addition there is a large overlap with solid state physics and materials engineering However solid state chemistry has still to be recognized by the general body of chemists as a legitimate subfield of chemistry The discipline is not even well defined as to content and has many facets that make writing a textbook a formidable task The early studies carried out in the United States by Roland Ward and his co workers emphasized the synthesis of new materials and the determination of their structure His work on doped alkaline earth sulfides formed the basis for the development of infrared phosphors and his pioneering studies on oxides were important in understanding the structural features of both the perovskite oxides as well as the magnetoplumbites In 1945 A F Wells published the first edition of Structural Inorganic Chemistry This work attempts to demonstrate that the synthesis structure and properties of solids form an important part of inorganic chemistry Now after almost 50 years during which many notable advances have been made in solid state chemistry it is still evident that the synthesis structure determination and properties of solids receive little attention in most treatments of inorganic chemistry The development of the field since the early studies of Roland Ward early 1940s has been rapid

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