

# Kinetics, Equilibrium and Thermodynamics: Practice Test

## 1. Multiple Choice

1. The primary rate equation is given by:  $k[A]^m[B]^n$  where:

- gradient of the graph is  $k[A]^m[B]^n$
- gradient of the graph is  $k[A]^m[B]^n$
- gradient represents all orders that contribute to the rate of reaction
- graph the concentration of substances over time to get order of reaction
- does not support

2. A study of reaction order with:

- by determining the half-life of reaction
- by plotting the rate of reaction
- by determining the activation energy of the reaction
- by plotting concentration against time (zero-order reaction only)
- order of the reaction is the order of the rate

3. Equilibrium following reaction:



At equilibrium,  $[\text{N}_2\text{O}_4] = 0.04 \text{ mol dm}^{-3}$  and  $[\text{NO}_2] = 0.08 \text{ mol dm}^{-3}$

The equilibrium constant is:

- 0.04
- 0.08
- 0.02
- 0.04
- 0.08

4. Equilibrium following reaction:



At equilibrium,  $[\text{N}_2] = 0.04 \text{ mol dm}^{-3}$  and  $[\text{H}_2] = 0.08 \text{ mol dm}^{-3}$

The equilibrium constant is:

- 0.04
- 0.08
- 0.02
- 0.04
- 0.08

5. The reaction mixture is at equilibrium:

- the reaction is at equilibrium
- the reaction is at equilibrium at the moment of observation
- the reaction is at equilibrium
- the reaction is at equilibrium (the reaction is at equilibrium at the moment of observation)
- the reaction is at equilibrium



6. The reaction mixture is at equilibrium:



At equilibrium, the reaction is at equilibrium:

- The temperature is high enough to provide enough energy to overcome the activation energy
- The temperature is high enough that the reaction mixture is at equilibrium
- The reaction is at equilibrium at the moment of observation
- The reaction is at equilibrium at the moment of observation

# Kinetics And Equilibrium Practice Test

**Neil D. Jespersen, Pamela Kerrigan**



## Kinetics And Equilibrium Practice Test:

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