

KINETICS & EQUILIBRIUM QUICK PRACTICE SHEETS

8.1 COLLISION THEORY

Worksheet: Work from left to right by following the path of correct answers. You do not mark your path.

Q1 Adding a catalyst

1.1 The reaction rate increases

1.2 The direction of change in which particles collide

1.3 One liter of a 2.0 M solution will react faster than 1 L of a 1.0 M solution

1.4 Increasing the pressure of a gas system will cause the reaction to proceed faster

1.5 Double gases readily react

1.6 A catalyst provides

1.7 Heat acts as a template for reactants and increases reaction rate

1.8 Increase pressure in gas systems for speed up reactions

1.9 A collision means the particles

1.10 Collision Theory explains how particles

8.5 EQUILIBRIUM

Worksheet: Complete the table and answer the questions.

Time (s)	Reactants (M)	Products (M)
0	1.00	0.00
10	0.80	0.20
20	0.72	0.28
30	0.65	0.35
40	0.60	0.40
50	0.55	0.45
60	0.52	0.48
70	0.50	0.50
80	0.50	0.50
90	0.50	0.50
100	0.50	0.50

$H_2 + I_2 \rightleftharpoons 2HI$

Questions:

- How does the concentration of the reactants and products change over time?
- When does the reaction approach equilibrium?
- When does the system reach equilibrium? How do you know?
- Compare the total concentration of the reactants and products.

8.6 LE CHATELIER'S PRINCIPLES

Worksheet: For each process, determine the reaction conditions, state signs in the table.

$NH_4Cl_{(s)} \rightleftharpoons NH_{3(g)} + HCl_{(g)} \quad \Delta H = +176 \text{ kJ/mol}$

1. Add ammonium chloride
2. Remove ammonium chloride
3. Add ammonia
4. Add hydrochloric acid
5. Remove ammonia
6. Remove hydrochloric acid
7. Increase temperature
8. Add a noble gas
9. Decrease temperature
10. Add a catalyst

6 WORKSHEETS

QUICK & EFFECTIVE



Kinetics And Equilibrium Mark Rosengarten Homework Answer

Arturo Cuomo



Kinetics And Equilibrium Mark Rosengarten Homework Answer:

Unveiling the Power of Verbal Beauty: An Psychological Sojourn through **Kinetics And Equilibrium Mark Rosengarten Homework Answer**

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