

- Which event must *always* occur for a chemical reaction to take place?
 - formation of a precipitate
 - formation of a gas
 - effective collisions between reacting particles
 - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
 - decreases
 - increases
 - remains the same
 - increases
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
 - a zinc strip and 1.0 M HCl(aq)
 - a zinc strip and 3.0 M HCl(aq)
 - zinc powder and 1.0 M HCl(aq)
 - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
 - 1.0 M HCl(aq) at 20.°C
 - 1.0 M HCl(aq) at 40.°C
 - 2.0 M HCl(aq) at 20.°C
 - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
 - lump
 - bar
 - powdered
 - sheet metal
- Given the balanced equation representing a reaction:

$$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$$

Which set of reaction conditions produces $\text{H}_2\text{(g)}$ at the fastest rate?

 - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
 - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
 - providing an alternate reaction pathway that has a higher activation energy
 - providing an alternate reaction pathway that has a lower activation energy
 - using the same reaction pathway and increasing the activation energy
 - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:



Which statement is true about energy in this reaction?

- The reaction is exothermic because it releases heat.
 - The reaction is exothermic because it absorbs heat.
 - The reaction is endothermic because it releases heat.
 - The reaction is endothermic because it absorbs heat.
- Which balanced equation represents an endothermic reaction?
 - $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
 - $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO(g)}$
 - Given the balanced equation:



Which statement best describes this process?

- It is endothermic and entropy increases.
 - It is endothermic and entropy decreases.
 - It is exothermic and entropy increases.
 - It is exothermic and entropy decreases.
- Which change is exothermic?
 - freezing of water
 - melting of iron
 - vaporization of ethanol
 - sublimation of iodine
 - Which reaction releases the greatest amount of energy per 2 moles of product?
 - $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
 - $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$
 - $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:

$$\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)} + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

 - It is exothermic and ΔH equals -91.8 kJ.
 - It is exothermic and ΔH equals +91.8 kJ.
 - It is endothermic and ΔH equals -91.8 kJ.
 - It is endothermic and ΔH equals +91.8 kJ.
 - Which expression represents the ΔH for a chemical reaction in terms of the potential energy, PE , of its products and reactants?
 - PE of products + PE of reactants
 - PE of products - PE of reactants
 - PE of products $\times$ PE of reactants
 - PE of products $\div$ PE of reactants

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DJ Losen



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web reynolds studied the conditions under which one type of flow changes into the other and found that the critical velocity at which laminar flow changes into turbulent flow depends on four quantities the diameter of the tube and the viscosity density and average linear velocity the liquid

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web fluid flow operation mccabe smith 3 3 used the lewis randal fugacity rule and the popular symbol for the same quantity
this edition only uses the popular notation and new problems have been added to the text finally the second edition includes
an appendix about the bridgman table and its use