CHEMISTRY 104 - Help Sheet #7

Chemical Kinetics-Part II

Do the topics appropriate for your course

Prepared by Dr. Tony Jacob

http://www.chem.wisc.edu/areas/ele (Resource page)

Nuggets: Evaluating Rate Laws: 3. Integrated Rate Laws, 4. Half-lives, 5. Rate versus Concentration, Integrated Rate Law Calculations; Half-Life

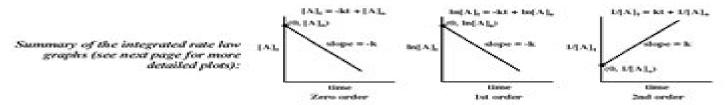
DIFFERENT METHODS TO DETERMINE RATE LAW (i.e., find the order(s) and k)

- I. Initial rates method (see Help Sheet #6)
- II. Description (see Help Sheet #6)
- III. Integrated Rate Laws Graphs (see below: this Help Sheet)
- IV. Half-life, t12: [A] versus time Graphs (see below: this Help Sheet)
- V. Rate versus Concentration Graphs (see below: this Help Sheet)
- VI. Mechanisms (see Help Sheet #8)

III. Integrated Rate Laws - Graphs and Equations

Integrated Rate Law: function of time; when a question involves time (e.g., How much [A] is left after 2 hours?) use the Integrated Rate Law equations

Onder	Rate Law	Integrated Rate Law	y-axis vs x-axis (y = mx + b)	Slope	y-intercept	t _{1/2} (half-life)	Units of k
0	$rate = \left. k[A]^O \right. = k$	$[A]_{\mathfrak{t}} = -k\mathfrak{t} + [A]_{\mathfrak{0}}$	[A] _t vs t	m = -k	[A] _o	$t_{1/2} = \frac{[A]_0}{2k}$	$\frac{M}{s} = N6s^{-1}$
1	$rate = k[A]^{I}$	$\ln[A]_{\mathfrak{t}} = -k\mathfrak{t} + \ln[A]_{\mathfrak{o}}$	ln(A) _t vs t	m = -k	In[A]o	$t_{1/2} = \frac{0.693}{k}$	$\frac{1}{s} = s^{-1}$
2	$rate = k[A]^2$	$\frac{1}{\{A\}} = kt + \frac{1}{\{A\}_0}$	$\frac{1}{ A _t}$ vs t	m - k	I [Al _o	$t_{1/2} = \frac{1}{k[A]_o}$	$\frac{1}{Ms} = M^{-1}s^{-1}$



1 st order rate law: $\ln[A]_t = -kt + \ln[A]_0$ can also be written: $\ln\left(\frac{[A]_t}{[A]_0}\right) = -kt$

this rearranged equation is useful when given % remaining questions, the % remaining (not % consumed) is substituted for $\left(\frac{[A]_q}{[A]_n}\right)$: $\ln(\% \text{remaining}) = -kt$

KINETICS TECHNIQUE: "Flooding" – see the pages at the end of this Help Sheet for a description and example; understanding this technique and interpreting the data is aligned with the lab and is often asked

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