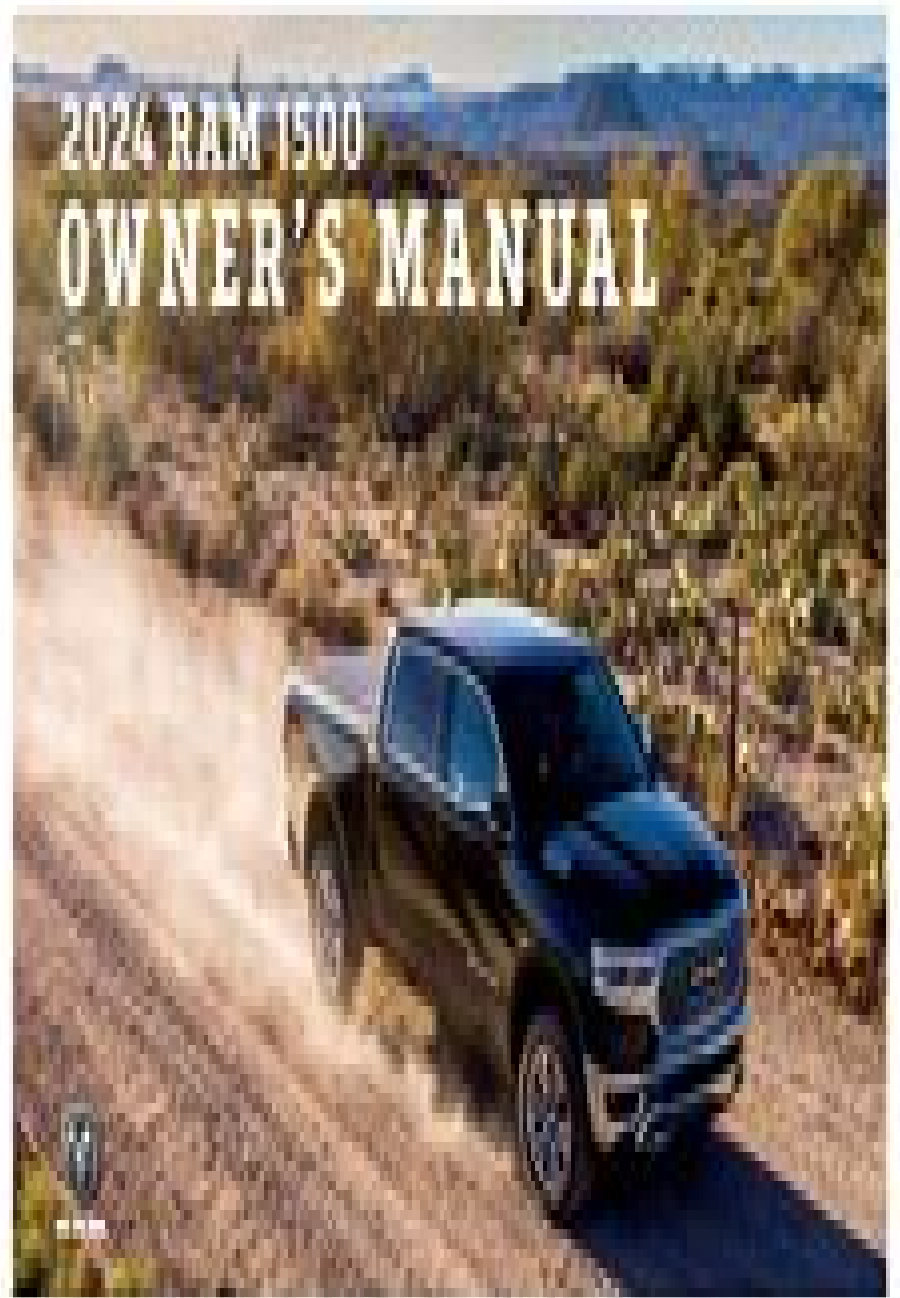


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web one of the ions produced is always h which immedi ately combines with a water molecule wo form che hyeonium ion h o the h o in is what defines the acidic properties of a substance because of thee special classification acids have a naming sytem dierent fom ianic or other molecular covalent compounds

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web acids are covalently bonded molecules but when they are put into water they produce ions one of the ions produced is

always H^+ which immediately combines with a water molecule to form the hydronium ion H_3O^+ the H_3O^+ ion is what defines the acidic properties of a substance

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web examine the pairs of ternary acids in model 2 that contain sulfur phosphorus and nitrogen each pair has one acid that ends in *ic* and another that ends in *ous* these endings are related to the name of the polyatomic anion found in the acid *ate* or *ite*

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web the acids lose a hydrogen ion for each acid base reaction in model 2 describe the role of the brønsted lowry base in the proton H^+ ion transfer that occurs the base gains or accepts a hydrogen ion proton as you saw in last week's pogil all arrhenius bases in model 1 have an OH^- ion in their chemical formulas

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web the molecular or empirical formula acids have their own set of naming rules cf sec 2 8 which you must also know you also need to be able to recognize and name a few organic compounds such as simple alkanes and alcohols see sec 2 9 key questions 8 in the chemical formula and name which element is given first a metal or nonmetal 9

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