IMPORTANT FORMULAE

1 SOLID STATE

- (1) Density (p) = $\frac{\alpha^3}{\alpha^3 N_A}$ [α^3 = Volume of unit cell
 - N_A = Avogadro number
 - M = Molar mass
 - n = No. of particles (atoms per unit cell)
- (4) Edge length and particle parameters in cubic system;

| (2) Mass | wiff. | wandt. | creft! | - | m | 100 | 41 |
|----------|-------|--------|--------|---|---|------------|----|
|----------|-------|--------|--------|---|---|------------|----|

where, m = mass of one particle)

Volume occupied by particles in unit cell

| of the last | When is builting on a page 1 to | | Section programming areas programming absolute | |
|-------------|---------------------------------|---|--|--|
| Sept. | Packing efficiency | - | Total volume of meit call × 100 | |

| Sr. No. | Unit Cell | Relation between | Volume of one particle | Total volume occupied by particles in unit cell |
|------------|----------------------------------|---|--|---|
| (9) | Simple cubic (se) | $r = \frac{a}{2} = 0.5000a$ | $\frac{\pi \alpha^3}{6} = 0.5237 \ \alpha^3$ | # a ³ - 0.5237 a ³ |
| OTTO | Body centred cubic lattice (bcc) | $r = \frac{\sqrt{3}}{4} \cdot a = 0.4330 \cdot a$ | $\frac{\sqrt{3} \times a^3}{16} = 0.34a^3$ | $\frac{\sqrt{3} \times a^3}{8} = 0.68a^3$ |
| trent | Face centred cubic lattice (foc) | $r = \frac{\sqrt{2}}{4} \cdot a = 0.3535 \cdot a$ | $\frac{\pi a^3}{12\sqrt{2}} = 0.185a^3$ | $\frac{\pi a^3}{3\sqrt{2}} = 0.74a^3$ |

050

(5) Coordination number and packing efficiency in system:

| Se. No. | Lattice | Co-ordination number | Packing efficiency |
|------------|-------------|-------------------------|-----------------------|
| 610 | BHC | 6 | 52.4% |
| pitty | bee | | 68% |
| (1001) | foc/eep/hep | 12 | 74% |

- (6) Number of particles and unit cells in metallic crystal:
 - (i) Number of particles in 'x' g = $\frac{xn}{p \cdot a^3}$
 - (iii) Number of unit cells in 'x' g metal = $\frac{x}{\rho \sigma^2}$
 - (iii) Number of unit cells in volume V of metal = $\frac{V}{a^3}$

2 SOURTIONS

- Henry's Law: S = P or S = K_HP where.
 - S = Solubility of the gas in mol L-1
 - P = Pressure of the gas in bar over the solution. K_{ss} = Henry's law constant
 - Unit of K_{el}: mol L⁻¹ = mol L⁻¹ bar⁻¹
- (2) Raoult's law (Vapour pressure of solutions of liquids in liquid) :
 - (8) $P_1 = x_1 P_1^{\circ}$ and $P_2 = x_2 P_2^{\circ}$...

where, P_1^0 and P_2^0 are vapour pressures of pure liquid A_1 and A_2 respectively.

 x_1 and x_2 = mole fraction of pure liquid A_1 and A_2 respectively.

(iii) Mole fraction of A = $\frac{\text{No. of moles of A}}{\text{No. of moles of solution}}$ = $\frac{n_A}{n_A + n_B}$

- (iii) Total Pressure $P = (P_0^0 P_1^0) x_0 + P_1^0$
- (iv) Composition of vapour phase: Composition of A and B in vapour are y, and y₂ respectively. P₁ and P₂ are the partial pressure of two components in the vapour, P = total vapour pressure.

 $P_1 = y_1P$, $P_2 = y_2P$ (as, $P_1 = P_1^0x_1$, $P_2 = P_2^0x_2$) where, y_1 and y_2 = mole fraction of two-components in the vapour.

$$\therefore y = \frac{x_1}{p} \text{ and } y_2 = 1 \cdot y_1$$

- (3) Colligative properties of non-electrolyte solution
- (A) Vapour pressure lowering:
 - (ii) Vapour pressure lowering = ΔP − P₁° − P₁ of P₁° is the vapour pressure of pure solvent and P₁ is the vapour pressure of solvent above the solution, P₁ < P₁°)
 - (iii) Raoult's law for solutions of monvolatile solutes $P_1 = P_1^{\circ}x_1 \quad (x_1 \text{ and } x_2 \text{ are the mole fraction}$ $\Delta P = P_1^{\circ}x_2 \quad \text{of solvent and solute respectively}$
 - (iii) Relative lowering of vapour pressure

$$x_0 = \frac{\Delta P}{P_1^0} = \frac{P_1^0 - P_1}{P_1^0}$$

(iv) Molar mass of solute from vapour pressure lowering

$$\frac{\Delta P}{P_1^0} = \frac{P_1^0 - P_1}{P_1^0} = \frac{W_2 M_1}{M_2 W_1}$$

[W₁ and W₂ = Mass of solvent and solute respectively (in grams) and M_1 and M_2 = Molar masses of solvent and solute[

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