

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + \text{Al}(\text{OH})_3(\text{s})$	None
2) $\text{NaCl}(\text{aq}) + \text{NH}_4\text{NO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$	None
3) $\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{AgI}(\text{s})$	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
4) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + (\text{Cu}_3(\text{PO}_4)_2(\text{s}))$	$3\text{Cu}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Cu}_3(\text{PO}_4)_2(\text{s})$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

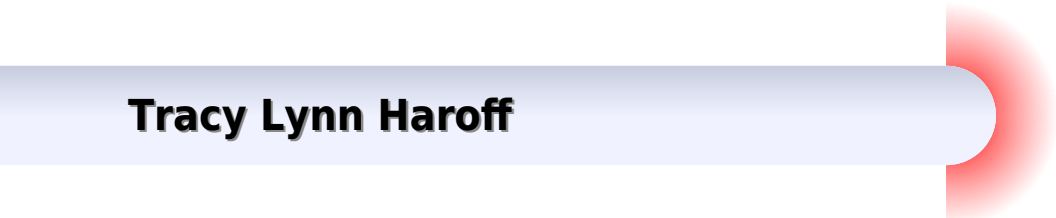
9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

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Tracy Lynn Haroff



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