

# Acid, Bases, pH and Buffers

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## Goals

- To measure the pH of some common household substances using cabbage juice indicator and pH meter.
- To observe the effect of added strong acid or strong base on the pH value of a buffer and unbuffered solution.

## Introduction

Plenty of household substances that we use in our daily life are either acidic or basic. Several theories have been proposed to theoretically explain the acidic and basic nature of chemicals.

*Arrhenius Theory:* According to the Arrhenius concept, acids are the substances that dissociates in water to produces hydrogen ions ( $H^+$ ), and bases are the substances that dissociates in water to produces hydroxide ions ( $OH^-$ ). Some common examples of Arrhenius acids are  $HNO_3$ ,  $HCl$ ,  $H_2SO_4$ , etc. and examples of Arrhenius bases are  $NaOH$ ,  $KOH$ ,  $Mg(OH)_2$  etc.



*Brønsted-Lowry (B-L) Theory:* According to the Brønsted-Lowry concept, acids are substances that can donate protons ( $H^+$ ) and bases are substances that can accept protons.



Everyday household products that contains acids in them include citrus fruits (lemons, oranges, etc.) and vinegar (acetic acid). The substances that contains bases in them are often household cleaners (soaps, bleach, detergents etc.).

Acid-base indicators, such as red cabbage juice or litmus paper, turn different colors depending on the pH and therefore can be used to test whether a solution is acidic or basic. Acids turn blue litmus paper red, and bases turn red litmus paper blue.

# Buffers In Household Products Prelab Answers

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