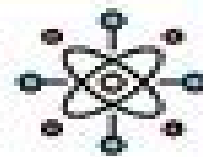


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by the attraction between positive metal ions and a sea of delocalized electrons. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding involves the sharing of electron pairs between atoms. This results in a stable molecule where the shared electrons are attracted to both nuclei. Covalent bonds are typically found in nonmetals and their compounds.

Ionic bonding involves the complete transfer of one or more electrons from one atom to another. This results in the formation of positively charged cations and negatively charged anions, which are held together by strong electrostatic forces.

Metallic bonding involves the attraction between positive metal ions and a sea of delocalized electrons. This results in a strong, malleable material that conducts electricity.

Science Real

Diagram illustrating the formation of ionic and covalent bonds.

Ionic Bonding: Shows a metal atom (e.g., Na) losing an electron to become a cation (Na^+) and a nonmetal atom (e.g., Cl) gaining an electron to become an anion (Cl^-). The resulting ionic bond is shown as the electrostatic attraction between these oppositely charged ions.

Covalent Bonding: Shows two nonmetal atoms (e.g., C and O) sharing pairs of electrons to form a covalent bond. The resulting molecule (e.g., CO_2) is shown with the shared electron pairs represented by lines or dots between the atoms.

Diagram illustrating the formation of a covalent bond: Two atoms, each with a valence shell, share a pair of electrons to form a covalent bond. The resulting molecule is shown with the shared electron pair represented by a line between the two atoms.

Diagram illustrating the formation of a metallic bond: A cluster of metal atoms is shown, with a sea of delocalized electrons surrounding them. The resulting metallic bond is shown as the attraction between the positive metal ions and the delocalized electrons.

Diagram illustrating the formation of a double bond: Two atoms share two pairs of electrons to form a double bond. The resulting molecule is shown with the shared electron pairs represented by two lines between the two atoms.

Diagram illustrating the formation of a triple bond: Two atoms share three pairs of electrons to form a triple bond. The resulting molecule is shown with the shared electron pairs represented by three lines between the two atoms.

GUIDED READING ACTIVITY

Key Covalent Bonds Guided Reading And Study

CL Gary



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