

- Which event must *always* occur for a chemical reaction to take place?
 - formation of a precipitate
 - formation of a gas
 - effective collisions between reacting particles
 - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
 - decreases
 - increases
 - remains the same
 - increases
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
 - a zinc strip and 1.0 M HCl(aq)
 - a zinc strip and 3.0 M HCl(aq)
 - zinc powder and 1.0 M HCl(aq)
 - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
 - 1.0 M HCl(aq) at 20.°C
 - 1.0 M HCl(aq) at 40.°C
 - 2.0 M HCl(aq) at 20.°C
 - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
 - lump
 - bar
 - powdered
 - sheet metal
- Given the balanced equation representing a reaction:

$$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$$

Which set of reaction conditions produces $\text{H}_2\text{(g)}$ at the fastest rate?

- a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
 - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
 - providing an alternate reaction pathway that has a higher activation energy
 - providing an alternate reaction pathway that has a lower activation energy
 - using the same reaction pathway and increasing the activation energy
 - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:



Which statement is true about energy in this reaction?

- The reaction is exothermic because it releases heat.
 - The reaction is exothermic because it absorbs heat.
 - The reaction is endothermic because it releases heat.
 - The reaction is endothermic because it absorbs heat.
- Which balanced equation represents an endothermic reaction?
 - $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
 - $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO(g)}$
 - Given the balanced equation:



Which statement best describes this process?

- It is endothermic and entropy increases.
 - It is endothermic and entropy decreases.
 - It is exothermic and entropy increases.
 - It is exothermic and entropy decreases.
- Which change is exothermic?
 - freezing of water
 - melting of iron
 - vaporization of ethanol
 - sublimation of iodine
 - Which reaction releases the greatest amount of energy per 2 moles of product?
 - $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
 - $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$
 - $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:

$$\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)} + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

 - It is exothermic and ΔH equals -91.8 kJ.
 - It is exothermic and ΔH equals +91.8 kJ.
 - It is endothermic and ΔH equals -91.8 kJ.
 - It is endothermic and ΔH equals +91.8 kJ.
 - Which expression represents the ΔH for a chemical reaction in terms of the potential energy, PE , of its products and reactants?
 - PE of products + PE of reactants
 - PE of products - PE of reactants
 - PE of products \times PE of reactants
 - PE of products \div PE of reactants

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Rates of Reaction and Equilibria Clifford Othen,1972-01-01 **Chemical Equilibrium** William Guenther,1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

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