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$\text{HNO}_3 + \text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{HC}_2\text{H}_3\text{O}_2$ chemical equation - Aug 11 2023

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web mar 4 2018 $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg}(\text{no}_3)_2 + \text{h}_2$ on the left side there are mg 1 magnesium in elemental form h_2 there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n_2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o_6

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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate mgco_3 and two moles of nitric acid hno_3 react to form one mole of magnesium nitrate $\text{mg}(\text{no}_3)_2$ one mole of water h_2o and one mole of carbon dioxide co_2 show chemical structure image reactants

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying hno_3 by 2 we get $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg}(\text{no}_3)_2 + \text{h}_2$ magnesium 2 nitric acid 2 magnesium nitrate 1 hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen

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