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Making Salts Worksheet

This worksheet accompanies *Making Salts Summary* and *Precipitation Reactions Summary*.

1. When an acid reacts with a base, a neutralisation reaction occurs, producing a salt. Using the following equation, name the salts that are produced in the acid-base reactions.

acid + base $\longrightarrow$ salt

- a) hydrochloric acid + ammonia $\longrightarrow$
- b) nitric acid + sodium hydroxide $\longrightarrow$
- c) phosphoric acid + potassium hydroxide $\longrightarrow$



2. There are three other ways of making salts from acids. Complete the following equations to illustrate them.

- a) acid + metal $\longrightarrow$ salt +
- b) acid + $\longrightarrow$ salt + water
- c) acid + carbonate $\longrightarrow$ salt + + water

3. Use the above information to complete the following reactions that form soluble salts using acids.

- a) iron oxide + sulfuric acid $\longrightarrow$ + water
- b) + hydrochloric acid $\longrightarrow$ magnesium chloride + H_2O + CO_2
- c) zinc + $\longrightarrow$ zinc nitrate + hydrogen

4. Name two ways in which salt is used by humans in everyday life.

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Acid Base Or Salt Physical Science If8767 Answers

**Jacques P. Fraissard, Leonidas
Petrakis, North Atlantic Treaty
Organization. Scientific Affairs
Division**

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Holt Science Spectrum Physical Science Chapter 9 Resource File: Acids, Bases, and Salts Holt Rinehart & Winston, 2008-01-01 Acids, Bases, and Salts Brian J. Knapp, 1998 **Acids, Bases, and Salts** Adams, This note is part of Quality testing Acids, Bases, and Salts William D. Adams (Childrens' author), 2023 Experience chemistry like never before The characters in this series will whisk you away on a colorful journey in science Do you get a bad reaction to chemistry Never fear Accelerate your learning with the Building Blocks of Chemistry Cover back **Acids and Bases** Brian G. Cox, 2013-01-31 Acids and bases are ubiquitous in chemistry Our understanding of them however is dominated by their behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described Acids, Bases & Salts , 2006

Acids, Bases and Salts Bill Freeman, 2003 Students explore the chemistry of acids and bases and focus on the pH scale a measurement of the amount of acidity and basicity of compounds Acidity and Basicity of Solids Jacques P. Fraissard, Leonidas Petrakis, North Atlantic Treaty Organization. Scientific Affairs Division, 1994 Historically technological developments that have made use of the acidity basicity of solids have often preceded an understanding of the phenomena involved This of course is very expensive and a far less efficient process than research based on a fundamental understanding of the science For the last 50 years therefore a vast amount of research has been devoted to the subject the rewards in terms

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