

CHEMISTRY TEST

NAME: _____

INTRODUCTORY CHEMISTRY CONCEPTS

DATE: _____

Directions: For each of the following questions, choose the number that **best** answers the question and place it on your answer sheet.

Directions: For each of the following questions, all major steps must be shown to receive full credit.

Introductory Chemistry Concepts**Level 1****Multiple Choice**

1. A chemical change always results characteristically in a change of

- 1) state; 2) form; 3) identity; 4) size; 5) shape.

ANS: 3

2. Which of the following is a chemical change?

- 1) evaporation of water 2) distillation of water 3) freezing of water
4) electrolysis of water

ANS: 4

3. Chemical changes always involve 1) the absorption of energy; 2) the liberation of energy;
3) either the absorption or liberation of energy; 4) changes of state.

ANS: 3

4. One physical property of sugar is its 1) acidity; 2) reactivity;

- 3) flammability; 4) plasma qualities; 5) solubility.

ANS: 5

5. The majority of chemical changes are 1) autothermic; 2) endothermic;

- 3) isothermic; 4) nonthermic; 5) exothermic.

ANS: 5

6. The statement which describes a **physical** property of the element oxygen is

- 1) oxygen supports the burning of paper; 2) oxygen is needed for the rusting of iron;
3) oxygen must be present for food to decay; 4) oxygen has a density of 1.43 grams per liter;
5) oxygen is a component of water.

ANS: 4

7. In any chemical reaction energy is 1) always absorbed; 2) always released;

- 3) released as heat; 4) usually not involved; 5) either absorbed or released.

ANS: 5

8. Which of the following is **not** a chemical change?

- 1) the dissolving of zinc by HCl 2) the exploding of gasoline 3) the corroding of a copper roof
4) the boiling of water

ANS: 4

9. One chemical property of a substance is a description of its 1) odor; 2) density;

- 3) combustibility; 4) solubility; 5) none of these.

ANS: 3

10. A new substance is formed as a result of a 1) physical change; 2) chemical

Chemistry Matter And Change Workbook Answer Key

CL Gary



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Table of Contents Chemistry Matter And Change Workbook Answer Key

1. Understanding the eBook Chemistry Matter And Change Workbook Answer Key
 - The Rise of Digital Reading Chemistry Matter And Change Workbook Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Workbook Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Workbook Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Workbook Answer Key
 - Personalized Recommendations
 - Chemistry Matter And Change Workbook Answer Key User Reviews and Ratings
 - Chemistry Matter And Change Workbook Answer Key and Bestseller Lists
5. Accessing Chemistry Matter And Change Workbook Answer Key Free and Paid eBooks
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 - Chemistry Matter And Change Workbook Answer Key eBook Subscription Services
 - Chemistry Matter And Change Workbook Answer Key Budget-Friendly Options
6. Navigating Chemistry Matter And Change Workbook Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Matter And Change Workbook Answer Key Compatibility with Devices
 - Chemistry Matter And Change Workbook Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Workbook Answer Key
 - Highlighting and Note-Taking Chemistry Matter And Change Workbook Answer Key
 - Interactive Elements Chemistry Matter And Change Workbook Answer Key
8. Staying Engaged with Chemistry Matter And Change Workbook Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Matter And Change Workbook Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Matter And Change Workbook Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Workbook Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Matter And Change Workbook Answer Key
 - Setting Reading Goals Chemistry Matter And Change Workbook Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Matter And Change Workbook Answer Key
 - Fact-Checking eBook Content of Chemistry Matter And Change Workbook Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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