

- Which of the following is/are not affected by temperature?
(a) Molarity (b) Molality
(c) Normality (d) None of these
- Ferric sulphate on heating gives sulphur trioxide. The ratio between the weights of oxygen and sulphur present in SO_3 obtained by heating 1 kg of ferric sulphate is
(a) 2 : 3 (b) 1 : 3
(c) 3 : 1 (d) 3 : 2
- The number of atoms present in 4.25 grams of NH_3 is approximately
(a) 1×10^{23} (b) 8×10^{28}
(c) 2×10^{23} (d) 6.02×10^{23}
- Two students performed the same experiment separately and each one of them recorded two readings of mass which are given below. Correct reading of mass is 3.0 g. On the basis of given data, mark the correct option out of the following statements:

Students	Readings	
	(i)	(ii)
A	3.01	2.99
B	3.05	2.95

- Results of both the students are neither accurate nor precise
- Results of student A are both precise and accurate
- Results of student B are neither precise nor accurate
- Results of student B are both precise and accurate
- What will be the molarity of a solution, which contains 5.85 g of NaCl (s) per 500 mL?
(a) 4 mol L^{-1} (b) 20 mol L^{-1}
(c) 0.2 mol L^{-1} (d) 2 mol L^{-1}
- Number of atoms in 55.85 gram Fe (at. wt. of Fe = 55.85 g mol^{-1}) is
(a) Twice that 60 g carbon (b) 6.023×10^{23}
(c) Half that in 8g He (d) $55.85 \times 6.023 \times 10^{23}$
- Neon has two isotopes Ne^{20} and Ne^{22} . If atomic weight of Neon is 20.2, the ratio of the relative abundances of the isotopes is
(a) 1 : 9 (b) 9 : 1
(c) 70 % (d) 80 %
- The total weight of 10^{22} molecular units of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ is nearly
(a) 4.144 g (b) 5.5 g
(c) 24.95 g (d) 41.45 g
- The number of Cl^- and Ca^{+2} ions in 222g. of CaCl_2 are
(a) $4N$, $2N$ (b) $2N$, $4N$
(c) $1N$, $2N$ (d) $2N$, $1N$

- The empirical formula of a gaseous compound is ' CH_2 '. The density of the compound is 1.25 gm/lit. at S.T.P. The molecular formula of the compound is 'X'
(a) C_2H_4 (b) C_3H_6 (c) C_6H_{12} (d) C_4H_8
- If 500 mL of a 5 M solution is diluted to 1500 mL, what will be the molarity of the solution obtained?
(a) 1.5 M (b) 1.66 M
(c) 0.017 M (d) 1.59 M
- The number of atoms present in one mole of an element is equal to Avogadro number. Which of the following element contains the greatest number of atoms?
(a) 4 g He (b) 46 g Na
(c) 0.40 g Ca (d) 12 g He
- The empirical formula of an organic compound is CH_2O . Its vapour density is 45. The molecular formula of the compound is
(a) CH_2O (b) $\text{C}_2\text{H}_4\text{O}_2$
(c) $\text{C}_3\text{H}_6\text{O}_3$ (d) $\text{C}_6\text{H}_{12}\text{O}_6$
- 0.132 g of an organic compound gave 50 ml of N_2 at NTP. The weight percentage of nitrogen in the compound is close to
(a) 15 (b) 20
(c) 48.9 (d) 47.34
- 0.7 moles of potassium sulphate is allowed to react with 0.9 moles of barium chloride in aqueous solutions. The number of moles of the substance precipitated in the reaction is
(a) 1.4 moles of potassium chloride
(b) 0.7 moles of barium sulphate
(c) 1.6 moles of potassium chloride
(d) 1.6 moles of barium sulphate
- The number of moles of Fe_2O_3 formed when 0.5 moles of O_2 and 0.5 moles of Fe are allowed to react are
(a) 0.25 (b) 0.5 (c) 1/3 (d) 0.125
- Amount of oxalic acid required to prepare 250ml of N/10 solution (MW of oxalic acid = 126) is
(a) 1.5759 g (b) 3.15 g (c) 15.75 g (d) 63.0 g
- The composition of compound A is 40% X and 60% Y. The composition of compound B is 25% X and 75% Y. According to the law of multiple Proportions the ratio of the weight of element Y in compounds A and B is:
(a) 1:2 (b) 2 : 1 (c) 2 : 3 (d) 3 :4
- If the concentration of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) in blood is 0.9 g L^{-1} , what will be the molarity of glucose in blood?
(a) 5 M (b) 50 M
(c) 0.005 M (d) 0.5 M
- What will be the molality of the solution containing 18.25 g of HCl gas in 500 g of water?
(a) 0.1 m (b) 10 m (c) 0.5 m (d) 1 m

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