

Compound **Q** is a viscous liquid which is very soluble in water.
The M_r of **Q** is 90.0.

Three possible structures for **Q** are shown below.

R	S	T
$\text{HOCH}_2\text{CH}_2\text{CO}_2\text{H}$	$\text{HOCH}_2\text{CO}_2\text{CH}_3$	$\text{HCO}_2\text{CH}_2\text{CH}_2\text{OH}$

(a) (i) What type of isomerism do **R**, **S** and **T** show?

.....

(ii) What oxygen-containing functional groups are present in **R**, **S** and **T**?
Give their **full names**.

R and

S and

T and

(iii) Which functional group(s) in (ii) will react with sodium carbonate?

.....

(iv) Which functional group(s) in (ii) will react with sodium metal?

.....

[6]

(b) When 0.002 mol of **Q** is reacted with an excess of solid sodium carbonate, Na_2CO_3 ,
24 cm³ of carbon dioxide, measured at room temperature and pressure, is produced.

(i) Calculate the amount, in moles, of carbon dioxide produced in this reaction.

(ii) Hence calculate the amount, in moles, of carbon dioxide produced by 1 mol of **Q**.

[2]

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
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
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