

3 • Molecules and Compounds**M O L A R M A S S & % C O M P O S I T I O N****I. Molar Masses**

Given a periodic table, you should be able to calculate the molecular mass (in u's) or the molar mass (in grams) for any element or compound.

Examples: (give answers to two decimal places)

H_2SO_4 98.09 amu	Cl_2 70.90 amu	Ca(OH)_2 74.10 amu	$\text{HC}_2\text{H}_3\text{O}_2$ 60.06 amu
CO_2 44.01 amu	N_2O 44.02 amu	NaOCl 74.44 amu	Al_2S_3 150.17 amu

II. Fraction and Percent Composition

It is useful to determine how much of a compound's mass is made up of each element. Water, H_2O , for example has a molar mass of 18.02 g. The H's mass is $2(1.0079) = 2.02$ g. The O's mass is 16.00 g.

We can set up fractions for each element: $\text{H} = \frac{2.02}{18.02} = 0.112 = 11.2\%$. $\text{O} = \frac{16.00}{18.02} = 0.888 = 88.8\%$.

This is called the **percent composition**. The fraction composition is a good in-between step.

Determine the fraction and percent composition of each element below (answer to one decimal place):

1. H_2SO_4	$\text{H} = \frac{2.02}{98.09} = 2.1\%$	$\text{S} = \frac{32.07}{98.09} = 32.7\%$	$\text{O} = \frac{64.00}{98.09} = 65.2\%$
2. Ca(OH)_2	$\text{Ca} = \frac{40.08}{74.10} = 54.1\%$	$\text{O} = \frac{32.00}{74.10} = 43.2\%$	$\text{H} = \frac{2.02}{74.10} = 2.7\%$
3. $\text{HC}_2\text{H}_3\text{O}_2$	$\text{H} = \frac{4.04}{60.06} = 6.7\%$	$\text{C} = \frac{24.02}{60.06} = 40.0\%$	$\text{O} = \frac{32.00}{60.06} = 53.3\%$
4. CO_2	$\text{C} = \frac{12.01}{44.01} = 27.3\%$	$\text{O} = \frac{32.00}{44.01} = 72.7\%$	
5. N_2O	$\text{N} = \frac{28.02}{44.02} = 63.7\%$	$\text{O} = \frac{16.00}{44.02} = 36.3\%$	
6. NaOCl	$\text{Na} = \frac{22.99}{74.44} = 30.9\%$	$\text{O} = \frac{16.00}{74.44} = 21.5\%$	$\text{Cl} = \frac{35.45}{74.44} = 47.6\%$
7. Al_2S_3	$\text{Al} = \frac{53.94}{150.17} = 35.9\%$	$\text{S} = \frac{96.21}{150.17} = 64.1\%$	

Chemistry Percent Composition Answers

Arshad Iqbal



Chemistry Percent Composition Answers:

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