LAB: THE PERIODIC LAW

Dmitri Mendeleev in 1869. The amazing accuracy of his predictions has been very important to chemists in this century. However, the basis of his arrangement was the atomic mastes of the elements. This approach proved incorrect as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the basis of the atomic numbers of the elements. In accordance with Moseley's revision, the periodic law states: the properties of the elements are periodic functions of their atomic numbers. Each of the 114 known elements has its own set of characteristic properties which include state, luster, melting and boiling points, electrical conductivity, reactivity, and so on. The elements are arranged within the periodic table into groups or families and periods or rows. This arrangement reflects the periodic or repeating nature of the properties of the elements. In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table. You will also predict values for any information missing from the table.

OBJECTIVES: In this exercise you will:

- arrange the elements in Groups 1, 2, 13-18 according to a list of clues and your knowledge of periodic properties.
- predict the missing properties of each element based on the location in the periodic table, and
- explain the trends of properties in families and periods.

EQUIPMENT:

Scissors Glue Element Clues Grids
Periodic table grid Scrambled element cards Element Clues sheet

PROCEDURE:

- *Obtain your <u>Element Clues Sheet</u> and <u>Element Clues Grids</u>. The code letters A to Z have been assigned to the first 26 main group elements in the periodic table. They do NOT correlate in any way with the real symbols for elements. Study the given clues, based on experimental data. Then place the code letters in their correct positions in Table A and fill in the correct atomic number and symbol for each in Table B.
- Obtain the <u>Scrambled Element Cards</u> and the blank periodic table grid. Cut out the first 26 squares corresponding to the first 26 main group elements. Position them in the blank periodic table grid according to the positions determined in step 1. Glue them in place; make sure they do not fall off.
- Cut out the remaining 16 blocks from the <u>Scrambled Element Cards</u> set. Use the information provided in
 each block and your knowledge of periodic trends to arrange these main group elements into their proper
 positions in the blank periodic table grid. When you are satisfied with the positions, glue the blocks into place.
- 4. Some of the information is missing from each of the <u>Scrambled Element Cards</u> blocks. Predict the values for the missing items from the location on the periodic table and write your prediction in the table. (Do NOT look up the values and write in the actual values; use your predictive skills and predict. Actual values will be considered WRONG and will be marked as such.)
- 5. Answer the conclusion questions for your lab.

CONCLUSION QUESTIONS:

- Examine your completed tale. What general observations can be made of trends within rows and within groups for the following properties: (a) density, (b) atomic radii, (c) melting point?
- Where are the heavy metals located? Cite three examples of heavy metals.
- 3. List four physical properties that distinguish metals from nonmetals.
- List the reason for the location of sodium in the periodic table. Give a least 3 reasons.
- 5. Explain the relationship of oxidation numbers to electron configuration for groups 1, 2, 3-18. How can an atom's electron configuration be predicted on the basis of its location in the periodic table?

TO TURN IN:

Do NOT turn in this direction sheet.

- Your completed Periodic Table Grid (<u>Building a Periodic Table</u>) with all cards securely glued to your paper.
- Your completed <u>Element Clues Grids</u>.
- Your answers to the conclusion questions on a separate sheet of paper, neatly written in complete sentences, attached to your periodic table. Include a proper lab heading.

Chemistry Periodic Law Lab Answers

H Kauffman

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