

Limiting Reagent and Percentage Yield

- 1 Consider the following equation:



- a Suppose 1.855 g of Cl_2 and 3.205 g of KBr are mixed in a solution. Which is the limiting reagent?

- b How many grams of each product are formed?

- 2 Suppose 25 L of nitrogen reacts with 25 L of hydrogen at STP to produce ammonia.



- a Identify the limiting and excess reagents. How many liters of ammonia gas are formed?

- 3 What is the percentage yield of a reaction in which 201 g of solid phosphorous trichloride reacts with excess water to form 128 g of aqueous hydrogen chloride and aqueous phosphorous acid? First, balance the chemical reaction.



Chemistry Theoretical And Percent Yield Answers

Thomas Griffiths



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