

Thermochemistry Review Questions

- The process of dissolving ammonium nitrate, NH_4NO_3 , in water is an endothermic process. What is the sign of 'q'? If you were to add some ammonium nitrate to water in a flask, would you expect the flask to feel warm or cold?
- Nitric acid, a source of many nitrogen compounds, is produced from nitrogen dioxide. An old process for making nitrogen dioxide employed nitrogen and oxygen: $N_{2(g)} + 2O_{2(g)} \rightarrow 2NO_{2(g)}$. The reaction **absorbs** 56.2 kJ of heat per 2 mol of $NO_{2(g)}$ produced. Is the reaction endothermic or exothermic? What is the value of q?
- Ammonia burns in the presence of a copper catalyst to form nitrogen gas: $NH_{3(g)} + 3O_{2(g)} \rightarrow 2N_{2(g)} + 6H_2O_{(g)}$ with a $\Delta H = -1267 \text{ kJ/mol}$. What is the enthalpy change (ΔH) to burn 35.8 g of ammonia? **[Ans: -2 663.45 kJ]**
- Propane, $C_3H_{8(g)}$, is a common fuel gas. Use the following to calculate the grams of propane you would need to provide 369 kJ of heat: $C_3H_{8(g)} + 5O_{2(g)} \rightarrow 3CO_{2(g)} + 4H_2O_{(g)}$ with a $\Delta H = -2043 \text{ kJ/mol}$. **[Ans: 7.96 g]**
- Ammonia burns in the presence of a platinum catalyst to form nitric oxide: $4NH_{3(g)} + 5O_{2(g)} \rightarrow 4NO_{(g)} + 6H_2O_{(g)}$. What is the heat of reaction at constant pressure? Use the following thermochemical equations **[Ans: - 906.3 kJ]**:



- Hydrogen sulfide, H_2S , is a poisonous gas with the odour of rotten eggs. The reaction for the formation of H_2S from the element is: $H_{2(g)} + \frac{1}{8}S_{8(rhombic)} \rightarrow H_2S_{(g)}$. Use Hess' law to obtain the enthalpy change for this reaction from the following enthalpies **[Ans: - 21 kJ]**:



- The cooling effect of alcohol on the skin is due to its evaporation. Calculate the heat of vaporization (ΔH) of ethanol, $C_2H_5OH_{(l)} \rightarrow C_2H_5OH_{(g)}$. The standard enthalpy of $C_2H_5OH_{(l)}$ is - 277.7 kJ/mol and that of $C_2H_5OH_{(g)}$ is - 235.4 kJ/mol. **[Ans: 42.3 kJ]**
- Carbon disulfide is a colourless liquid. When pure, it is nearly odourless, but the commercial produce smells vile. Carbon disulfide is used in the manufacture of rayon and cellophane. The liquid burns as follows: $CS_{2(l)} + 3O_{2(g)} \rightarrow CO_{2(g)} + 2SO_{2(g)}$. Calculate the standard enthalpy change for this reaction using standard enthalpies of formation **[Ans: - 1076.6 kJ]**.
- The group IIA carbonates decompose when heated. For example: $BaCO_{3(s)} \rightarrow BaO_{(s)} + CO_{2(g)}$. Use the following enthalpies of formation and calculate the heat require to decompose 10 g of barium carbonate. The standard enthalpy of $BaCO_3$ is - 1216.3 kJ/mol. $BaO_{(s)}$ is - 942.3 kJ/mol and $CO_{2(g)}$ is - 393.5 kJ/mol. **[Ans: - 6.09 kJ]**.

Chemistry Thermochemistry Review Questions

Julia Burdge



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