

30. 1.4 What is the pH of a solution whose pOH = 12.6? $pH + pOH = 14$
 $14 - 12.6$
31. 1.0×10^{-14} What is the product of H_3O^+ and OH^- concentrations in water?
32. 1.0×10^{-7} What is the concentration of H_3O^+ in pure water?
33. If the $[H_3O^+]$ in a solution is $2.5 \times 10^{-3} M$, and the $[OH^-]$ is $4.0 \times 10^{-6} M$, the solution (acidic/basic/neutral). 0.000000025 0.0000040

basic; the $[OH^-]$ is greater.

34. Calculate the $[OH^-]$ of a solution whose $[H_3O^+] = 5.2 \times 10^{-5}$ (ans. ~~1.9×10^{-9}~~)

$$\frac{[H_3O^+][OH^-]}{[H_3O^+]} = \frac{1.0 \times 10^{-14}}{5.2 \times 10^{-5}}$$

$$2 \times 10^{-9}$$

35. If $[H_3O^+] = 2.7 \times 10^{-4} M$, what is the pH of the solution? (ans. 3.6)

$$pH = -\log(2.7 \times 10^{-4})$$

$$pH = 3.57$$

36. If $[OH^-] = 2.35 \times 10^{-9} M$, what is the pOH of the solution? (ans. 8.63)

$$pOH = -\log(2.35 \times 10^{-9})$$

$$pOH = 8.63$$

37. If the $[OH^-]$ is known to be 3.6×10^{-9} , what is the $[H_3O^+]$? (ans. 2.8×10^{-6})

$$\frac{[H_3O^+][OH^-]}{[OH^-]} = \frac{1.0 \times 10^{-14}}{[OH^-]}$$

$$[H_3O^+] = 2.78 \times 10^{-6}$$

38. Is the solution in the problem above (acidic, basic or neutral)?

acidic

because $[H_3O^+]$ is greater

39. How many moles of NO_2 are present in 2.41×10^{24} molecules? (ans. 4.00 mol)

$$\frac{2.41 \times 10^{24} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules/mol}} = 4.00 \text{ mol}$$

40. For the combustion reaction of methane: $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$. How many moles of CO_2 are produced if you start with 164 g of CH_4 ? (ans. 10.2 mol)

$$\frac{164 \text{ g } CH_4}{16.05 \text{ g } CH_4} \times \frac{1 \text{ mol } CH_4}{1 \text{ mol } CH_4} \times \frac{1 \text{ mol } CO_2}{1 \text{ mol } CH_4} = 10.22 \text{ mol}$$

molar mass

mole : mole ratio (coefficients)

Chemistry Unit 11 Review Acids And Bases Answers

G Psacharopoulos



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