

9. From the following list of substances, circle the ones that are elements:

silver	carbon dioxide	wood	alcohol
water	hydrogen	carbon	nitrogen
oxygen	gold	sugar	salt
air	sulfur	magnesium	nickel

10. Classify the following as pure substances or as mixtures:

air M	gasoline P.S.	water P.S.	sugar P.S.
mercury P.S.	oxygen P.S.	salt water M	gold P.S.

11. Classify the following as heterogeneous or as homogeneous:

sand & salt mixture	bronze (includes copper and tin) H O
salt water	air H E
tossed salad H E	iron with rust H E
	granite H O

12. Provide an example of an element and a compound.

Hydrogen

water

13. Provide examples of how mixtures can be separated. Explain when each type of technique should be used.

Filtration → 1 dissolves in water + 1 doesn't

Distillation → 2 liquids with different boiling points

magnet → 1 is magnetic + 1 isn't

14. Explain the difference between physical and chemical change and provide example for both.

Physical change - retain the same substance afterwards

↳ Boiling point

Chemical - creates a new substance

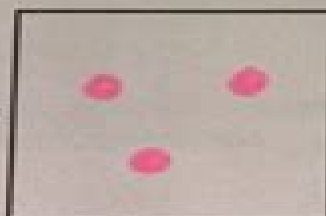
↳ combustion

15. Suppose that one volume of gas A combined with two volumes of gas B to form one volume of product when measured at the same pressure and temperature. Sketch particle diagrams for molecules of gas A, gas B and the product; assume gases A and B are monatomic.

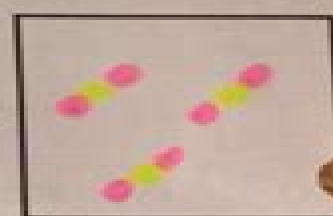
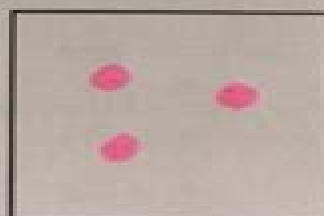


gas A

+



gas B



product

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**S Nieto**



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