

1) Define the following terms:

Atomic Number:

Mass Number:

Atomic Mass:

Isotope:

Isotope Name:

Nuclear Symbol:

*another side*2) Give the nuclear symbol for an atom with 9 protons and 10 neutrons: $^{19}_{9}\text{F}$ 3) Give the nuclear symbol for an atom with atomic number 72 and 100 neutrons: $^{172}_{72}\text{Hf}$ 4) Write the isotope name for an element with 20 protons and 21 neutrons: $^{41}_{20}\text{Ca}$

5) Complete this chart for the following NEUTRAL atoms using your periodic table:

Element	Atomic #	Protons	Electrons	Neutrons	Mass Number	Isotope Name	nuclear symbol
Si	14	14	14	15	29	Si-29	$^{29}_{14}\text{Si}$
B	5	5	5	5	10	B-10	$^{10}_{5}\text{B}$
F	9	9	9	10	19	F-19	$^{19}_{9}\text{F}$
U	92	92	92	143	235	U-235	$^{235}_{92}\text{U}$
Kr	36	36	36	48	84	Kr-84	$^{84}_{36}\text{Kr}$
U	92	92	92	146	238	U-238	$^{238}_{92}\text{U}$
Xe	54	54	54	77	131		$^{131}_{54}\text{Xe}$

6) Which two of the above would be an example of isotopes?

U-235 & U-238

7) Explain the difference between "atomic mass" and "mass number"

mass number refers to a specific isotope's # of p⁺ + n⁰
atomic mass is an average of all isotopic masses

8) Complete this chart:

Particle	Charge	Mass in amu	Location in Atom	Discovered by?
proton	$+$	1	nucleus	Rutherford
neutron	0	1	nucleus	Chadwick
electron	$-$	$\frac{1}{1836}$	e ⁻ cloud	Thomson

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JR Anderson



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