

QUIZZZ

AP Chemistry Unit 8 Review
81 questions / 24
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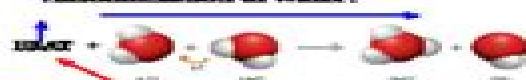
Unit 8

Acids & Bases

2.

8.1 Introduction to Acids & Bases

Autoionization of Water:



This is an endothermic process- increasing the heat will shift the reaction to produce more H_3O^+ and OH^- . K_w increases (more products), both pH and pOH will decrease but pure H₂O will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$$\text{pH} = 3.45$$

is sig. figs

In case of acidic solution $> [\text{H}^+] > [\text{OH}^-]$

$$\text{pH} < 7.00 \text{ at } 25^\circ\text{C}$$

In case of neutral solution $> [\text{H}^+] = [\text{OH}^-]$

$$\text{pH} = 7.00 \text{ at } 25^\circ\text{C}$$

In case of basic solution $> [\text{H}^+] < [\text{OH}^-]$

$$\text{pH} > 7.00 \text{ at } 25^\circ\text{C}$$

ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$K_w = K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also known:

$$[\text{H}^+] = 10^{-\text{pH}} \text{ and } [\text{OH}^-] = 10^{-\text{pOH}}$$

3. The value of K_w for water at 0°C is 1×10^{-15} . What is the pOH of water at 0°C ?

a) 7.5

b) 8.0

c) 7.0

d) 6.5

Chemistry Unit 8 Review V2

Chao Zhang



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Table of Contents Chemistry Unit 8 Review V2

1. Understanding the eBook Chemistry Unit 8 Review V2
 - The Rise of Digital Reading Chemistry Unit 8 Review V2
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 8 Review V2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 8 Review V2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 8 Review V2
 - Personalized Recommendations
 - Chemistry Unit 8 Review V2 User Reviews and Ratings
 - Chemistry Unit 8 Review V2 and Bestseller Lists

5. Accessing Chemistry Unit 8 Review V2 Free and Paid eBooks
 - Chemistry Unit 8 Review V2 Public Domain eBooks
 - Chemistry Unit 8 Review V2 eBook Subscription Services
 - Chemistry Unit 8 Review V2 Budget-Friendly Options
6. Navigating Chemistry Unit 8 Review V2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 8 Review V2 Compatibility with Devices
 - Chemistry Unit 8 Review V2 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 8 Review V2
 - Highlighting and Note-Taking Chemistry Unit 8 Review V2
 - Interactive Elements Chemistry Unit 8 Review V2
8. Staying Engaged with Chemistry Unit 8 Review V2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 8 Review V2
9. Balancing eBooks and Physical Books Chemistry Unit 8 Review V2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 8 Review V2
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit 8 Review V2
 - Setting Reading Goals Chemistry Unit 8 Review V2
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit 8 Review V2
 - Fact-Checking eBook Content of Chemistry Unit 8 Review V2
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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