

# Covalent Bonding

## How does a covalent bond form?

- A **covalent bond** forms when two atoms **share** one or more pairs of electrons.
- Atoms in covalent bonds typically have **similar electronegativities** and are **usually non-metals**.
- The shared electrons allow each atom to achieve a full outer electron shell, similar to the noble gas configuration.

## Types of Covalent Bonds

- **Single Bond:** One pair of electrons is shared (e.g.,  $\text{H}_2$ ,  $\text{Cl}_2$ ).
- **Double Bond:** Two pairs of electrons are shared (e.g.,  $\text{O}_2$ ,  $\text{CO}_2$ ).
- **Triple Bond:** Three pairs of electrons are shared (e.g.,  $\text{N}_2$ ).



## Dot-and-Cross Diagrams

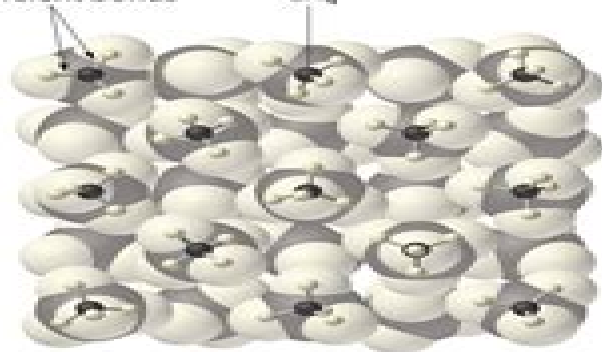
- Only the outer electron shells (valence electrons) are shown.
- Shared pairs of electrons between atoms are shown as pairs of dots and crosses.
- It represents how the atoms bond and achieve stable electron configurations.

## Example: Water $\text{H}_2\text{O}$



Covalent bonds

$\text{CH}_4$



## Properties of Covalent Compounds

- **Molecular Structure:** They are made up of discrete molecules with **weak intermolecular forces**.
- **Melting and Boiling Points:** Generally **lower** than ionic compounds because the forces between molecules are weaker.
- **Electrical Conductivity:** **Poor conductors** of electricity in all states, as they do not have free ions or electrons.
- **Solubility:** Often **soluble in nonpolar solvents** and may be insoluble in water, depending on their polarity.

# Chemistry Study Guide Covalent Bonding

**Camilla Rothe**



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