

Chemistry I Final Review

Determine the amount of heat absorbed or released in each of the following changes:

1. 25.0 g of water heated from 15.0°C to 35.0°C . (cal)

2. 25.0 g of water cooled from 35.0°C to 20.0°C . (cal)

3. 25.0 g of water heated from 35.0°C to 45.0°C . (J)

4. 115.0 g of water cooled from 35.0°C to 25.0°C . (J)

5. How many grams of C_2H_2 must be burned in an excess supply of oxygen to produce 11.4 MJ of heat energy?



6. Which of the following reactions are exothermic?

a. $\text{H}_2\text{O}(g) + 7.4\text{ kJ} \rightarrow \text{H}_2\text{O}(l)$ and $44.5\text{ kJ} + \text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$

b. $30\text{ kJ} + \text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$ and $110\text{ kJ} + \text{H}_2\text{O}(g) \rightarrow \text{H}_2\text{O}(l)$

c. $\text{H}_2\text{O}(g) \rightarrow \text{H}_2\text{O}(l)$ and $\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$

d. $\text{H}_2\text{O}(g) + 20\text{ kJ} \rightarrow \text{H}_2\text{O}(l)$ and $20\text{ kJ} + \text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$

7. Predict the heat of reaction for the following reaction:



(Use the tables for any data required in this problem.)

Chemistry Thermochemistry Unit Test Review

DP Hallahan



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