

Name: _____

Date: _____

Mole Concept Questions Part 1 (No concentrations and limiting agents)

(Please show workings neatly on a clean piece of paper)

- 1) Calculate the percentage composition of hydrogen in ammonium chloride. [2m]
- 2) What does "the mole" represents or means? [1m]
- 3) How many atoms are there in 1 mole of helium gas? [1m]
- 4) How many molecules are there in 1 mole of carbon dioxide gas? [1m]
- 5) Calculate the number of atoms in 0.005 mol of carbon dioxide? [2m]
- 6) Calculate the number of molecules of NaCl in a packet of NaCl weighing 150g. [2m]
- 7) Calculate the number of atoms in 7dm³ of NO₂. [2m]
- 8) Determine the empirical formulae of a compound containing 87g of copper and 48g of oxygen. [3m]
- 9) Calculate the percentage yield of the following reaction when 25g of NaCl is produced when 1 mole of Na₂CO₃ reacts with excess HCl.

$$2\text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$$
- 10) We have an impure calcium oxide weighing 5g reacting with sulfuric acid. Given that 7.5g of calcium sulfate is produced, which is the percentage purity of the calcium oxide in the sample?

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