

Chapter Goals 1.1

Intro. to Chemical Reactions

Name: _____
 Date: _____
 Class: _____

Electrolyte Review

1. Write the formula for the following compounds:

- a) barium phosphate: $\text{Ba}_3(\text{PO}_4)_2$ b) sodium phosphate: Na_3PO_4
 c) magnesium phosphate: $\text{Mg}_3(\text{PO}_4)_2$ d) lithium nitrate: LiNO_3

2. Write the name for the following compounds:

- a) $\text{Cu}(\text{NO}_3)_2$: copper(II) nitrate b) FeCl_3 : iron(III) chloride
 c) SnSO_4 : tin(II) sulfate d) SnF_4 : tin(IV) fluoride

Chemical Reactions: What they look like...

Here is what a chemical reaction looks like when it is written:



There are examples that you can write to describe what happens above:

Chemical reactions result with atoms and molecules and electron and energy.

There are also some characteristics for chemical reactions and the reactions themselves. Here are a few points:

Critical Thinking Questions

1. Write the following reaction, with the reactants and products:



Chemical reactions result with atoms and molecules and electron and energy.

Chemquest Answer Key

Rachel Turoscy



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The Directory of Management Consultants, 2003 Kennedy Information (Firm), 2002

The Answer Key: A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success (Volume One) (First Edition) Rachel Turoscy, 2018-08-09 The Answer Key A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success Volume 1 is a concise and accessible textbook that covers the critical information a student needs to understand the basic mathematics used in chemistry courses The book provides easy to understand step by step instructions for solving general chemistry problems The book begins with chapters dedicated to problem solving methodology and unit conversions In subsequent chapters the text covers important topics like ionic and covalent bonding chemical formula calculations solubility and reactions in aqueous solution gases the first law of Thermodynamics Quantum theory and electron configuration It also covers periodic trends the Lewis Dot Structures and bonding theories Each chapter contains sample problems and practice problems to help further understanding of how math and chemistry go hand in hand The Answer Key is an excellent resource for any undergraduate course that deals with the basic concepts of general chemistry

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