

Chapter 7 Chemical Reactions

Section 7.1 Describing Reactions

(pages 192–198)

This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.

Reading Strategy (page 192)

Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

Chemical Equations (pages 192–193)

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$.
 - Carbon and oxygen react and form carbon monoxide.
 - ☒ Carbon and oxygen react and form carbon dioxide.
 - Carbon dioxide yields carbon and oxygen.
 - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
 - 1
 - 7
 - ☒ 14
 - 28
- In the reaction represented by the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

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Jianjun Gao



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Table of Contents Chemistry Prentice Hall Answer Key Chemical Reactions

1. Understanding the eBook Chemistry Prentice Hall Answer Key Chemical Reactions
 - The Rise of Digital Reading Chemistry Prentice Hall Answer Key Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Prentice Hall Answer Key Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Prentice Hall Answer Key Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Prentice Hall Answer Key Chemical Reactions

- Personalized Recommendations
 - Chemistry Prentice Hall Answer Key Chemical Reactions User Reviews and Ratings
 - Chemistry Prentice Hall Answer Key Chemical Reactions and Bestseller Lists
5. Accessing Chemistry Prentice Hall Answer Key Chemical Reactions Free and Paid eBooks
 - Chemistry Prentice Hall Answer Key Chemical Reactions Public Domain eBooks
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 - Chemistry Prentice Hall Answer Key Chemical Reactions Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chemistry Prentice Hall Answer Key Chemical Reactions Compatibility with Devices
 - Chemistry Prentice Hall Answer Key Chemical Reactions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Prentice Hall Answer Key Chemical Reactions
 - Highlighting and Note-Taking Chemistry Prentice Hall Answer Key Chemical Reactions
 - Interactive Elements Chemistry Prentice Hall Answer Key Chemical Reactions
 8. Staying Engaged with Chemistry Prentice Hall Answer Key Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Prentice Hall Answer Key Chemical Reactions
 9. Balancing eBooks and Physical Books Chemistry Prentice Hall Answer Key Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Prentice Hall Answer Key Chemical Reactions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Prentice Hall Answer Key Chemical Reactions
 - Setting Reading Goals Chemistry Prentice Hall Answer Key Chemical Reactions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Prentice Hall Answer Key Chemical Reactions

- Fact-Checking eBook Content of Chemistry Prentice Hall Answer Key Chemical Reactions
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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