

2-



When a 2.000-gram sample of pure phenol, $\text{C}_6\text{H}_5\text{OH}(l)$, is completely burned according to the equation above, 64.98 kilojoules of heat is released. Use the information in the table below to answer the questions that follow.

Substance	Standard Heat of Formation, ΔH_f° , at 25°C (kJ/mol)	Absolute Entropy, S° , at 25°C (J/mol-K)
C(graphite)	0.00	5.69
$\text{CO}_2(g)$	-393.5	213.6
$\text{H}_2(g)$	0.00	130.6
$\text{H}_2\text{O}(l)$	-285.85	69.91
$\text{O}_2(g)$	0.00	205.0
$\text{C}_6\text{H}_5\text{OH}(l)$?	144.0

- (a) Calculate the molar heat of combustion of phenol in kilojoules per mole at 25°C. -3058 kJ
- (b) Calculate the standard heat of formation, ΔH_f° , of phenol in kilojoules per mole at 25°C. -161 kJ
- (c) Calculate the value of the standard free-energy change, ΔG° , for the combustion of phenol at 25°C. -3032 kJ
- (d) If the volume of the combustion container is 10.0 liters, calculate the final pressure in the container when the temperature is changed to 110.°C. (Assume no oxygen remains unreacted and that all products are gaseous.) $.601 \text{ atm}$

Chemistry Unit 10 Review Answers

**Avril Higton, Mike Clemmet, Elaine
Golding, Alan V Jones**



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