

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

Chemistry Unit 15 Worksheet Answer

L Darling-Hammond



Chemistry Unit 15 Worksheet Answer:

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Table of Contents Chemistry Unit 15 Worksheet Answer

1. Understanding the eBook Chemistry Unit 15 Worksheet Answer
 - The Rise of Digital Reading Chemistry Unit 15 Worksheet Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 15 Worksheet Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 15 Worksheet Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 15 Worksheet Answer
 - Personalized Recommendations
 - Chemistry Unit 15 Worksheet Answer User Reviews and Ratings

- Chemistry Unit 15 Worksheet Answer and Bestseller Lists
- 5. Accessing Chemistry Unit 15 Worksheet Answer Free and Paid eBooks
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 - Chemistry Unit 15 Worksheet Answer eBook Subscription Services
 - Chemistry Unit 15 Worksheet Answer Budget-Friendly Options
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 - Chemistry Unit 15 Worksheet Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 15 Worksheet Answer
 - Highlighting and Note-Taking Chemistry Unit 15 Worksheet Answer
 - Interactive Elements Chemistry Unit 15 Worksheet Answer
- 8. Staying Engaged with Chemistry Unit 15 Worksheet Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 15 Worksheet Answer
- 9. Balancing eBooks and Physical Books Chemistry Unit 15 Worksheet Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 15 Worksheet Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Unit 15 Worksheet Answer
 - Setting Reading Goals Chemistry Unit 15 Worksheet Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Unit 15 Worksheet Answer
 - Fact-Checking eBook Content of Chemistry Unit 15 Worksheet Answer
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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